

Characterisation and applications of synthesised cation exchanger guar gum sulphonic acid (GSA) resin for removal and recovery of toxic metal ions from industrial wastewater

AV Singh* and Naresh Kumar Sharma

Applied and Environmental Laboratory, Department of Chemistry, Jai Narain Vyas University, Jodhpur – 342033, (Rajasthan) India

Abstract

The chemically modified guar gum sulphonic acid (GSA) resin was used for removal and pre-concentration of Zn^{2+} , Cd^{2+} , Fe^{2+} , Pb^{2+} and Cu^{2+} ions in aqueous solutions and steel industry effluent from Jodhpur, India. This type of ion exchange resin represents a new class of hybrid ion exchangers with higher stability, reproducibility and selectivity for toxic metal ions. The characterisation of the resin was carried out by determining the ion exchange capacity, elemental analysis, FT-IR spectra and thermal analysis (TGA). The distribution coefficients (K_d) of toxic metal ions were determined for different pH values. The adsorption of different metal ions on GSA resin follows the order: $Cu^{2+} > Fe^{2+} > Zn^{2+} > Cd^{2+} > Pb^{2+}$. The adsorbency of different metal ions on the GSA resin was studied for up to 10 cycles.

Keywords: distribution coefficients, Ion exchange capacity, GSA resin, toxic metal ions

Introduction

Water pollution caused by toxic metal ions from various industrial effluents has attracted great attention worldwide. A number of regulatory agencies have endeavoured to restrict the discharge of untreated metal-containing effluents into public sewers, rivers, the sea and onto land. The removal of toxic metal ions is very important due to their toxic nature in water, posing serious hazards to human health as well as to other living organisms. Therefore, it is essential to improve methods for the removal of metal ions to decrease the pollution load on the environment.

Different methods, such as precipitation (Harper and Kingham, 1992), solvent extraction (Kiezyk and Mackay, 2009; Rios et al., 2010) chemical and electrochemical techniques (Diasz et al., 2009), advanced oxidation processes (Kepa et al., 2008; Diwani et al., 2009), ion-exchange methods (Saha et al., 2000; Wang et al., 2002; Singh and Saraf, 2009; Das et al., 1999; Lata et al., 2003), ultrafiltration and reverse osmosis (Rengaraj et al., 2001; Yurlova et al., 2002; Benito and Ruiz, 2002) have been developed for the removal of toxic metal ions from industrial effluents and wastewater. A major drawback with precipitation is sludge production. Adsorption using commercial activated carbon (CAC) can remove toxic metal ions from wastewater, such as Cd (Ramos et al., 1997), Ni (Shim et al., 2001), Cr (Ouki and Neufeld, 1999) and Cu (Monser and Adhoum, 2002). However, CAC remains an expensive material for toxic metal removal. The resins prepared with DVB-styrene backbone are hydrophobic (Sarin et al., 1980). Also these resins are prepared from petrochemicals, which are very costly. The ion exchange capacity of these resins depends mainly on the quantity of functional groups and the pH of the solution. The most widespread

chelating functional groups are: thiol (Kantipuly et al., 2001), thiourea (Hubicki and Leszczynska, 2009), dithizone (Chwastowska and Kosiarska, 2001; Shah and Devi, 1998; Chwastowska et al., 2004) and triisobutyl phosphine sulfide (Hidalgo et al., 2001; Sanchez et al., 2001), which are used for the removal of toxic metal ions from effluent; slow adsorption kinetics is one of their undesired characteristics.

Ion-exchangers of polyvalent metals often exhibit much better ion-exchange behaviour and are always of interest because of their higher selectivity, resistance to higher temperatures, and to some extent chemical stability, than those of commercial ion-exchange resins (Vesely and Pekark, 2001; Inoue, 2003). Inorganic ion-exchange materials have also been established to have applications in various disciplines, i.e. environmental studies, medical science (kidney dialysis), ion-selective electrode preparation (Hassan et al., 2000), heterogeneous solid-state membrane formation, pre-concentration and ion-exchange fibre preparation. Inorganic-organic materials are known to have a large selectivity for mono- and multivalent cations (Maya, 2003; Clearfield, 2001; Khan and Mezboul Alam, 2003; Khan et al., 2004; Amicangelo and Leenstra, 2005; Patel et al., 2006).

The increasing number of publications on adsorption of toxic compounds by modified polysaccharides indicates a recent increase in interest in the synthesis of new low-cost adsorbents used in wastewater treatment (Crini, 2005). Recent developments have focused on the synthesis of adsorbents containing polysaccharides, in particular modified biopolymers derived from chitin (Santose et al., 2007), chitosan (Wan Nagh et al., 2002), starch (Crini et al., 2005), cellulose (El-Mensawy et al., 2008; Gupta et al., 2002) and cyclodextrin (Kozlowski et al., 2005), which are not only eco-friendly and cost-effective but are also effective in remediation of common effluents present in the wastewater.

Guar gum is a polysaccharide-galactomannan obtained from the seeds of the leguminous arid plant *Cymopsis tetragonolobus*. The ratio of mannose to galactose is 2:1. The structure of guar gum is shown in Fig. 1.

* To whom all correspondence should be addressed.

☎ +91 9461286995; fax: +91 0291 2720436;

e-mail: nareshksharma89@rediffmail.com

Received 25 October 2010; accepted in revised form 30 May 2011.

Chauhan et al. (2009) studied the synthesis and characterisation of novel guar gum hydrogels and their use as Cu(II) sorbent (Chauhan et al., 2009). Synthesis of a guaran-sulphonic acid cation-exchanger and its application in metal ion removal from underground mine-water of the Rajpura Dariba Mines, Udaipur, India, was reported by Singh et al. (1996). A hydrophilic polysaccharide matrix of guaran has been used for the preparation of some new chelating resins, after its cross linking with epoxychloropropane (Ahuja and Rai, 1997).

The ion exchangers based on guar gum powder are hydrophilic and biodegradable, whereas other ion exchangers are prepared from petrochemical products which are not hydrophilic and biodegradable. Due to rising prices of petroleum products the guar gum powder has been selected for development of guar gum sulphonic acid (GSA) resin, the cost of which is very low, because it is locally available in large quantities from agriculture resources. These biopolymers are also environmentally friendly.

In this work, the synthesis and characterisation of GSA resin and its applications for removal and recovery of toxic metal ions from reference solution and effluent of Apex Steel Industry, Jodhpur, India, is described. The effects of pH, agitation speed, treatment time, flow rate, contact time, temperature and adsorbent dose, on the removal of metal ions from industrial effluent, were investigated.

Materials and methods

Chemicals and sample

Materials used were as follows: guar gum powder (Ases Chemical Works, Jodhpur, India); epoxychloropropane (Aldrich, USA); sulphonic acid (Loba Chemic Pvt Ltd, Mumbai, India), sodium hydroxide (Sarabhai M. Chemicals, Baroda, India), 1,4-dioxacyclohexane (E Merk, Mumbai, India). The reference solution and the effluent of Apex Steel Industry, Jodhpur (Rajasthan) characteristic features are presented in Tables 1 and 2, respectively. The reference solution was used to observe the removal efficiency of GSA resin in the experiment.

Metal ion	Fe ²⁺	Cu ²⁺	Zn ²⁺	Pb ²⁺	Cd ²⁺
Concentration (in mg·l ⁻¹)	45	28	40	25	55

Appearance	: Turbid						
Colour	: Green						
pH	: 4.5						
Total hardness	: 966 mg·l ⁻¹						
Metal ion	Fe ²⁺	Cu ²⁺	Zn ²⁺	Pb ²⁺	Cd ²⁺	Mg ²⁺	Ca ²⁺
Concentration (in mg·l ⁻¹)	1.13	2.02	9.28	0.98	0.18	19.20	80.0
fluoride = 0.27 ; sulphate = 840.12 ; cyanide = 0.04							

Methods

Perkin-Elmer Model 2380 atomic absorption spectrometer (AAS) operating with an air-acetylene flame was used to

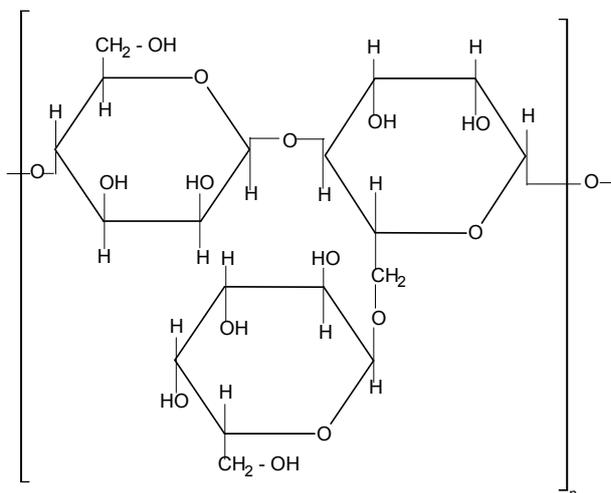


Figure 1
Structure of guar gum

analyse the concentration of metal ions in solutions. The metal ion concentration in the aqueous solution was determined by the AAS technique. Varian FTIR (Fourier transform infrared spectroscopy) instrument was used for determination of infrared (IR) spectrum using KBr pellets. FTIR spectral analysis of the resin was done in H⁺ form. Thermogravimetric analyser Model 951 was used for determination of the thermal stability of synthesised resin. Carlo Erba Model 1160 elemental analyser was used for determination of CHSO in the synthesised resin.

Synthesis of guar gum sulphonic acid (GSA) resin

To prepare the cation exchange resin based on guar gum powder, it was treated with sodium 1,2-epoxypropyl sulphonate, which was prepared by the reaction of epoxychloropropane with sodium bisulphite. This compound was previously synthesised by Singh et al. (1996).

Swelling studies

The H⁺ form of dry resin (0.6 g) was equilibrated with 60 ml of distilled water for 3 days. The swollen resin was collected by filtration; adhering traces of water were removed by filter-pressing and the sample was weighed. The swollen resin was dried in a vacuum for 24 h and weighed again. The following equation was used to determine the equilibrium water calculated (EWC):

$$\text{EWC} = \frac{\text{Weight of wet resin} - \text{Weight of dry resin}}{\text{Weight of wet resin}}$$

Column experiment

In the column experiment, a glass tube with 1.6 cm internal diameter and 20 cm height, packed with 9 cm of resin (8.5 g), was used. 50 ml of the reference metal ion solution was passed through the column at a flow rate of 2 ml·min⁻¹. The flow rate was controlled by a peristaltic pump. The column was washed with 20 ml of deionised water and the washing was rejected. The metal ions were eluted quantitatively with different strengths of acids.

Determination of distribution coefficient

The distribution coefficient (K_d) of metal ions on resins was determined by the batch method. In all cases, for the determination of K_d a 50 ml reference solution was taken in a conical flask and the pH was adjusted by acetic acid + sodium acetate and ammonium hydroxide + ammonium chloride. 20 mg of GSA resin were added to the solution and stirred on a magnetic stirrer for 2 h and the contents were equilibrated. The solution was filtered through Whatman No. 40 filter paper. The residue on the filter paper was equilibrated with 4 N HCl, and the solution was filtered through Whatman No. 42 filter paper.

The toxic metal ion concentrations in the filtrate as well as in the residue were estimated using atomic absorption spectrophotometry. The calibration curves for different metal ions were plotted, by analysing a series of standard solutions of metal ions using AAS. The concentration of metal ions in filtrate was determined by the calibration curves. The percentage removal of the metal ions and their distribution coefficient (K_d) on GSA resin were calculated using the following formula (Khan et al., 2009):

$$K_d = \frac{\text{Amount of metal ion in resin phase}}{\text{Amount of metal ion in solution phase}} \times \frac{\text{Volume of solution (ml)}}{\text{Weight of dry resin (g)}}$$

$$\text{i.e. } K_d = [(I - F) / F] \times V / M \text{ (ml} \cdot \text{g}^{-1}\text{)}$$

where:

- I is the initial amount of the metal ion in solution
- F is the final amount of metal ion in solution after equilibrium with resin
- V is the volume of metal ion solution (ml)
- M is the weight of the resin taken (g).

Determination of percentage removal of metal ions concentration

The initial metal ion concentrations in solution and filtrate after equilibrium with resin were estimated using an atomic absorption spectrophotometer. The percentage removal of metal ions was calculated using the following formula:

$$\text{Percentage removal of metal ions} = [(I - F) / I] \times 100$$

where:

- I = initial concentration of metal ion in solution;
- F = final concentration of metal ions in solution after equilibrium with resin.

Determination of ion exchange capacity (IEC)

Resin capacity is usually expressed in terms of equivalents per litre ($\text{eq} \cdot \text{L}^{-1}$) of resin or milliequivalents per dry gram of resin. The ion exchange capacity, which is generally taken as a measure of the hydrogen ion liberated by neutral salt to flow through the composite cation exchanger, was determined by standard column process. 1 g (dry mass) of the composite ion exchange material in H^+ form was placed in a glass column with a glass wool support at the bottom. It was washed with demineralised water to remove any excess acid remaining on the particles.

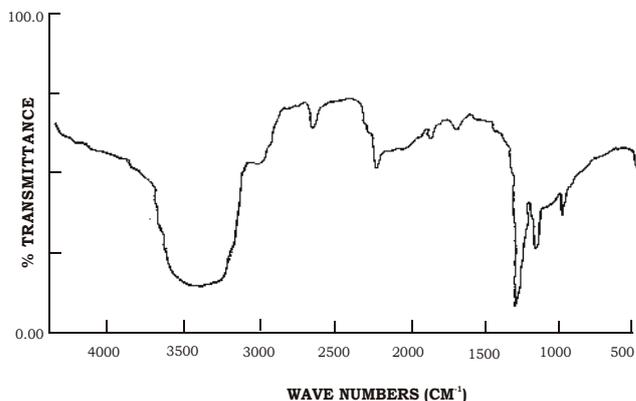


Figure 2
FTIR spectrum of guar gum sulphonic acid (GSA) resin

The hydrogen ions were eluted with 0.1 M solution of different alkali and alkaline earth salts. The flow rate was kept at $2 \text{ ml} \cdot \text{min}^{-1}$. The collected effluent was titrated against a standard solution of sodium hydroxide using phenolphthalein as an indicator. The hydrogen ions released were then calculated.

Results and discussion

Spectral characterisation of guar gum sulphonic acid (GSA) resin

FTIR spectrum of GSA resin shows a peak at 2950 cm^{-1} which is attributed to C-H stretching vibration. A strong peak in the region $1250\text{--}1000 \text{ cm}^{-1}$ denotes C-O stretching vibrations. The GSA resin displays asymmetric and symmetric S=O stretching frequencies in the region $1350\text{--}1342 \text{ cm}^{-1}$ and $1165\text{--}1150 \text{ cm}^{-1}$. The frequency of S=O absorption is increased by the electronegative substitution. Another variable peak at $1600\text{--}1500 \text{ cm}^{-1}$ is attributed to C-H bending. The spectra of polysaccharides are generally observed in the region $3600\text{--}3200 \text{ cm}^{-1}$, which denotes O-H stretching frequencies (Rao, 1963). The IR spectrum of GSA resin is given in Fig. 2.

Thermogravimetric analysis

The polymer sample was powdered to the same average mesh size and dried carefully in a vacuum desiccator. The boat was packed uniformly for analysis. For the dynamic measurement, the system was heated at a constant heating rate of $20^\circ\text{C} \cdot \text{min}^{-1}$ under a static atmosphere until complete decomposition. The GSA resin was found to be stable up to 401°C , after which degradation was found to be rapid. The TGA curve of GSA resin is shown in Fig. 3.

Elemental analysis

The results of the elemental analysis were in good agreement with calculated value. The elemental analysis results were in support of the proposed structure of the resin presented in Fig. 1. The results of elemental analysis are given in Table 3.

Element	C	H	S	O
Calculated %	49.61	5.94	7.24	37.21
Found %	48.50	5.05	6.55	36.89

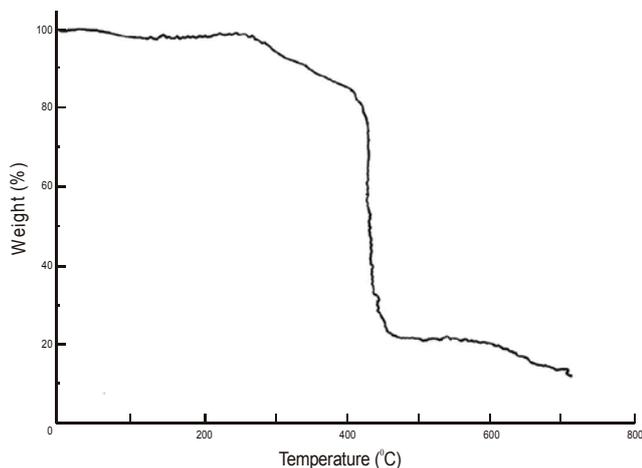


Figure 3
TGA curve of GSA resin

Swelling studies

Higher swelling of resins may create problems in column operation. For high swelling resins additional allowances need to be made in the column design. Higher swelling means the solvent can easily penetrate in the resin phase and solute adsorption rate increases. In Indion 225H resin swelling was 14% (Gwale and Marathe, 2007) and in GSA resin swelling was 6.11%.

Moisture contents

A physical property of the ion exchange resins which changes with cross linkage is the moisture content of the resin. The amount of cross linking depends on the proportions of different monomers used in the polymerisation step. It is thus important to determine the moisture contents of the synthesised resins. 1 g of the resin in hydrogen form was taken and dried to a constant weight in a vacuum desiccator, at 85°C overnight, and the resin was weighed. Moisture percentage was 14%.

Resin durability

It was observed that the adsorbency of different metal ions on the GSA resin after 10 cycles (adsorption and desorption), and ion exchange capacities of resin, were almost constant. The adsorbed metal ions were easily desorbed by treatment with different strengths of acids, at room temperature. Table 4 shows the adsorbency of different metal ions on the GSA resin after 10 cycles (adsorption and desorption).

Metal ions	Adsorption percentage for metal ions onto GSA resin after a number of cycles				
	1 cycle	2 cycles	4 cycles	8 cycles	10 cycles
Cu ²⁺	96.05	94.86	92.10	92.05	92.01
Fe ²⁺	95.88	93.51	91.47	91.42	91.36
Zn ²⁺	94.50	92.40	91.33	91.29	91.25
Cd ²⁺	93.08	92.98	91.82	91.70	91.68
Pb ²⁺	90.10	89.05	88.51	88.45	88.41

Ion exchange capacity (IEC) of GSA resin

The IEC of the dry GSA resin was found to be 9.50 meq·g⁻¹.

Distribution coefficient (K_d) of metal ions from effluent of Apex Steel Industry, Jodhpur

The pH has a strong effect on the distribution coefficient (K_d) of metal ions. The distribution coefficients (K_d) of metal ions from the effluent of Apex Steel Industry, Jodhpur, are given in Table 5. The results show that the distribution coefficient value first increases, and then decreases, with increasing pH, and that these optimum results were obtained at pH 7.0. Differences in distribution coefficient, at the same pH, for different metal ions suggest a possible strategy for separation of these ions from their mixtures. Dowex 50W synthetic resin shows different K_d values at the same pH for Cu²⁺, Zn²⁺, Ni²⁺, Cd²⁺, and Pb²⁺ metal ions (Pehlivan and Altun, 2006). Metal sorption starts when the pH rises to the range where most acidic ion exchange sites start to exchange hydronium ion for metal, and the capacity reaches the maximum value in the pH range where all of the ion exchange sites take part in the reaction and the functional group is able to form a complex with the metal cations (Nenov and Bonev, 2010).

pH	Zn ²⁺	Cd ²⁺	Fe ²⁺	Pb ²⁺	Cu ²⁺
3	3.05	87.50	60.60	51.56	27.06
4	7.58	87.51	131.94	86.36	194.56
5	16.72	87.51	131.95	111.11	227.50
6	35.73	200.00	231.81	163.46	290.62
7	421.15	425.00	540.00	220.00	606.25
8	180.30	87.50	17.80	39.47	14.45
9	56.69	65.00	10.31	29.44	9.12

Statistical analysis

The relative standard deviation values (RSD) of the optimum removal percentages of metal ions are shown in Table 6. All data represent the mean of 3 independent experiments.

Metal ions	Amount of metal ions in effluent (mg)	Amount loaded on GSA resin (mg)	Percentage removal	RSD %
Cu ²⁺	2.02	1.94	96.05	1.47
Fe ²⁺	1.13	1.08	95.88	0.70
Zn ²⁺	9.28	8.76	94.50	1.26
Cd ²⁺	0.18	0.16	93.08	1.77
Pb ²⁺	0.98	0.88	90.10	1.93

Recovery of metal ions

The metal ions were eluted quantitatively with different strengths of acids. Zn²⁺ was eluted with 0.05 N HCl; Fe²⁺ with 0.5 N HCl; Cu²⁺ with 1.5 N HCl; Pb²⁺ with 0.5 N HNO₃ and Cd²⁺ was eluted with 1 N HNO₃. The resin column was then washed thoroughly with demineralised water. These results are

reported in Table 7. Dev and Roar (1998) modified amberlite XAD-4 with bis-(N,N- salicylidene)-1,3-propanediamine for the separation of Cu^{2+} , Fe^{2+} , Ni^{2+} , Co^{2+} , Zn^{2+} , Hg^{2+} and Pb^{2+} , and sorbed metals were eluted by 1 M HCl with a recovery of 97 to 100%.

Metal ions	Amount loaded on GSA resin (mg)	Amount found (mg)	Percentage recovery	Eluent used	Eluent volume (mℓ)
Zn^{2+}	8.76	8.39	95.86	0.05 N HCl	50
Fe^{2+}	1.08	1.06	98.81	0.5 N HCl	55
Cu^{2+}	1.94	1.92	99.38	1.5 N HCl	50
Cd^{2+}	0.16	0.15	97.28	1 N HNO_3	45
Pb^{2+}	0.88	0.88	100.00	0.5 N HNO_3	40

Effect of pH

The pH is an important parameter for adsorption of metal ions from aqueous solution because it affects the solubility of the metal ions, concentration of the counter ions on the functional groups of the adsorbent and the degree of ionisation of the adsorbent during reaction. To examine the adsorption percentage of metal ions with pH, the pH was varied from 3.0 to 9.0. The uptake of free metal ions depends on pH; optimum adsorption of metal ions was observed at pH 7 and declined at higher pH. Adsorption of metal ions on GSA resin increased in the pH range from 3.0 to 7.0 (Table 8). The observed effect of pH on adsorption is due to different tendencies for chelate formation at different pH values. This can be related to differences in the stability of metal complexes formed on resin. Beyond pH 10.0, the insoluble metal ions start precipitating from the solution, making true sorption studies impossible.

pH	Zn^{2+}	Cd^{2+}	Fe^{2+}	Pb^{2+}	Cu^{2+}
3	10.97	79.30	71.37	67.36	52.27
4	23.34	80.02	83.87	77.46	88.44
5	40.17	80.53	84.07	81.30	89.91
6	58.88	87.73	90.73	86.34	91.91
7	94.50	93.08	95.88	90.10	96.05
8	87.84	78.98	41.39	61.54	36.68
9	69.50	77.76	29.48	54.57	26.61

Effect of agitation speed

The effect of agitation speed on adsorption of metal ions was studied by varying the speed of agitation from 0 (without shaking) to 180 $\text{r}\cdot\text{min}^{-1}$, while keeping the optimum temperature (25°C) and pH constant. It is clear from Fig. 4 that the adsorption of metal ions on ion exchange resin generally increased with increasing agitation speed. The adsorption of metal ions on GSA resin increased when agitation speed increased from 0 to 120 $\text{r}\cdot\text{min}^{-1}$. These results can be attributed to the fact that the increase in agitation speed improves the diffusion of metal ions towards the surface of the adsorbents. This also indicates that a shaking rate in the range 120-180 $\text{r}\cdot\text{min}^{-1}$ is sufficient to ensure that all of the surface binding sites are made readily available for metal ion uptake. The effect of external film

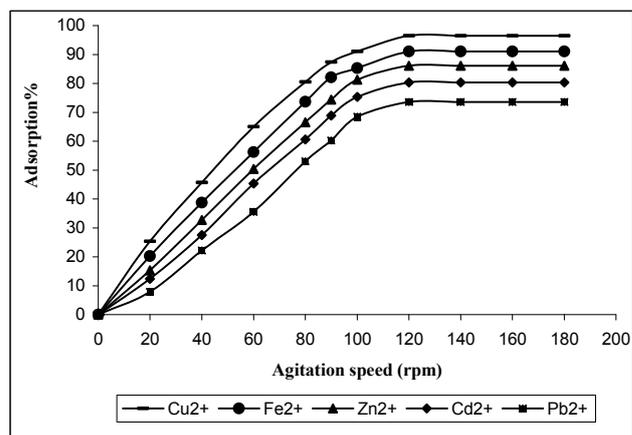


Figure 4
Effect of agitation speed on adsorption of metal ions by GSA resin

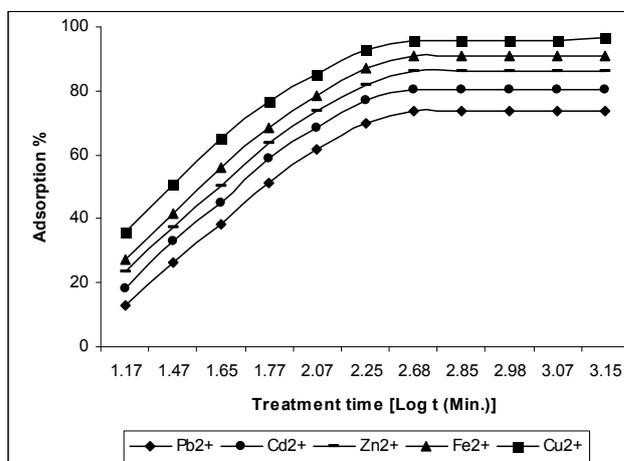


Figure 5
Effect of changing treatment time on the adsorption of different metal ions on GSA resin

diffusion on adsorption rate can then be assumed to not be significant. For convenience, agitation speed of 130 $\text{r}\cdot\text{min}^{-1}$ was selected as the optimum speed for GSA resin for removal of metal ions from effluent of Apex Steel Industry. These results are in close agreement with those reported by Jeon and Park (2005).

Effect of treatment time

The results for treatment time indicate that the adsorption of metal ions increased with an increase in contact time, before equilibrium is reached. It is clear from Fig. 5 that adsorption of metal ions on GSA resin increased when contact time was increased from 30 to 240 min; optimum contact time for GSA adsorbent was found to be 240 min. Other parameters, such as pH of solution and agitation speed were kept optimum, while temperature was kept at 25°C . Greater availability of sulphonic and hydroxyl functional groups on the surface of guar gum, which are required for interaction with metal ions, significantly improved the binding capacity and the process proceeded rapidly. This result is important, as equilibrium time is one of the important parameters for an economical wastewater treatment system.

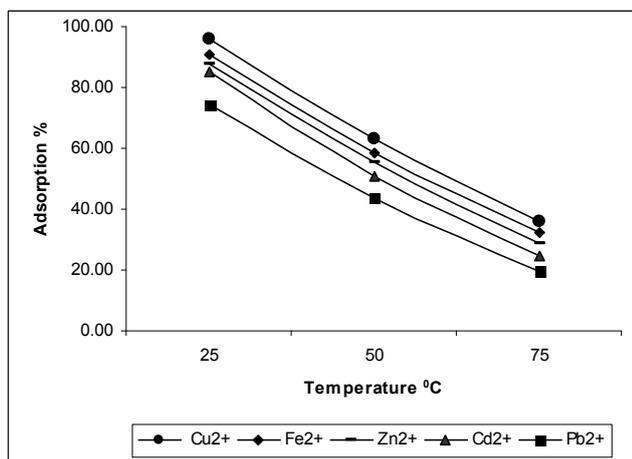


Figure 6
Effect of changing treatment temperature on the adsorption of different metal ions on GSA resin

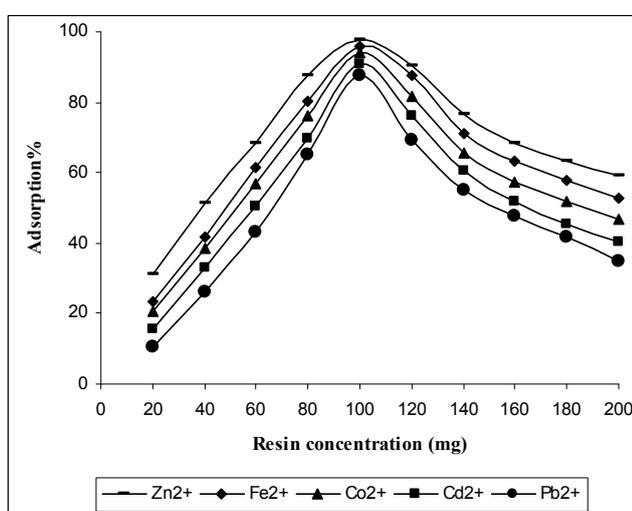


Figure 7
Effect of adsorbent dose on adsorption of metal ions on GSA resin

Effect of treatment temperature

Figure 6 shows the effect of treatment temperature on the adsorption percentage of the metal ions on GSA resin. The adsorption percentage of metal ions decreases by increasing the treatment temperature from 25°C to 75°C at optimum treatment time of 4 h. This is due to the exothermic ion exchange reactions of divalent cations. The equilibrium constants for such reactions decrease with temperature and the reaction products were not favoured at high temperatures. This observation is in full agreement with the published results of Khalil and Farag (1998).

Effect of GSA dose on removal of metal ions

The adsorption of metal ions was significantly influenced by the amount of GSA resin added. The amount of GSA added to the solution determined the number of binding sites available for adsorption. The effect of adsorbent dose on the amount of metal ion removed was studied by the application of varying GSA doses. The maximum adsorption by GSA was achieved

with an adsorbent dose of 0.1 g, after which adsorption decreased until a dose of 0.2 g. The initial increase in adsorption of metal ions was due to the availability of more adsorption sites. On increasing the GSA concentration further, the binding of metal ions was steadily decreased. This effect might be attributed to overlapping or aggregation of adsorption sites of resin, resulting in a decrease in the total surface area of the adsorbent. These results are shown in Fig. 7.

Effect of flow rate on removal of metal ions

In the column experiment, the flow rate of the sample solution was an important parameter, not only affecting the recovery of metal ions but also controlling the time of analysis. Therefore, the effect of flow rate on removal of metal ions was investigated under the optimum conditions (pH, eluent, etc.) by passing 100 ml of sample solution through the micro column. The flow rates were adjusted in the range of 1.0-5.0 ml·min⁻¹, controlled by a peristaltic pump. It was found that metal ions can be adsorbed quantitatively by GSA resin at a flow rate below 1.5 ml·min⁻¹. Above 5 ml·min⁻¹, the recovery was almost 93%. Flow rates lower than 1.0 ml·min⁻¹ were not employed, in order to avoid the long extraction time. Thus, a flow rate of 2.0 ml·min⁻¹ was selected for use throughout the column experiment.

Conclusion

The adsorption behaviour of GSA resin towards Cu²⁺, Fe²⁺, Zn²⁺, Cd²⁺ and Pb²⁺ ions from aqueous solution was investigated under different conditions, including agitation speed, pH, treatment time and temperature. It was found that GSA resin is an effective adsorbent (almost 96%) for the removal of different toxic metal ions from the effluent of Apex Steel Industry, Jodhpur, India, in the following sequence, Cu²⁺ > Fe²⁺ > Zn²⁺ > Cd²⁺ > Pb²⁺. It has been found that the maximum percentage removals for Cu²⁺, Fe²⁺, Zn²⁺, Cd²⁺ and Pb²⁺ are 96.05%, 95.88%, 94.50%, 93.08% and 90.10% respectively. The optimum results were obtained at pH 7.0. GSA resin is thus recommended for treatment of effluents containing the aforementioned metal ions.

Acknowledgment

The authors are thankful to the Head, Department of Chemistry, JNV University, Jodhpur, for providing all necessary facilities.

References

- AHUJA M and RAI AK (1997) Adsorption studies with some chelating ion exchange resins derived from guaran. *Carb. Polym.* **33** (1) 57-62.
- AMICANGELO JC and LEENSTRA WR (2005) Zirconium arene-phosphonates: Chemical and structural characterization of 2-naphthyl and 2-anthracenylphosphonate Systems. *Inorg. Chem.* **44** (6) 2067-2073.
- BENITO Y and RUIZ ML (2002) Reverse osmosis applied to metal finishing wastewater. *Desalination* **142** (3) 229-234.
- CHAUHAN K, CHAUHAN GS and AHN JH (2009) Synthesis and characterization of novel guar gum hydrogels and their use as Cu (II) sorbent. *Biosour. Technol.* **100** (3) 3599-3603.
- CHWASTOWSKA J and KOSIARSKA E (2001) Synthesis and analytical characterization of chelating resin loaded with dithizone. *Talanta* **35** (6) 439-442.

- CHWASTOWSKA J, SKWARA W, STERLIN SKA E and PSZONICKI L (2004) Determination of platinum and palladium in environmental samples by graphite furnace atomic absorption spectrometry after separation on dithizone sorbent. *Talanta* **64** (1) 224-229.
- CLEARFIELD A (2001) Group IV phosphates as catalysts and catalyst supports. *J. Mol. Catal.* **27** (1-2) 251-262.
- CRINIF G, BERTINI S, FILIATRE C and TORRI G (2005) Preparation, characterization and sorption properties of crosslinked starch-based exchangers. *Carbohydr. Polym.* **60** (1) 67-75.
- CRINI G (2005) Recent developments in polysaccharide-based materials used as adsorbents in waste water treatment. *Prog. Polym. Sci.* **30** (1) 38-70.
- DAS D, DAS AK and SINHA C (1999) A new resin containing benzimidazolylazo group and its use in the separation of heavy metals. *Talanta* **48** (5) 1013-1022.
- DEV K and ROAR GN (1998) Synthesis and analytical properties of a chelating resin functionalised with bis-(N, N'-salicylidene) 1, 3-propanediamine ligands. *Talanta* **43** (3) 451-457.
- DE LOS RIOS AP, HERNANDEZ-FERNANDEZ FJ, LOZANO LJ, SANCHEZ S, MORENO JI and GOLINEZ C (2010) Removal of metal ions from aqueous solutions by extraction with ionic liquids. *J. Chem. Eng. Data* **55** (2) 605-608.
- DIAZ CB, HERNANDEZ IL, MORALS GR, BILYEU B and HERNANDEZ PB (2009) Removal of biorefractory compounds in industrial wastewater by chemical and electrochemical pretreatments. *Ind. Eng. Chem. Res.* **48** (3) 1253-1258.
- DIWANI GEI, RAFIE SE and HAWASH S (2009) Use of sonication advanced oxidation process for purification of energetic waste water. *Am.-Euras. J. Agri. Environ. Sci.* **6** (1) 119-128.
- EL-MENSNAWY AM, SALAMA A and EL-ASMY AA (2008) New modified cellulose for separation and determination of mercury in environment water sample. *Can. J. Anal. Sci. Spec.* **53** (5) 206-213.
- GAWALE R and MARATHE KV (2007) Removal of metal ions from micellar solution. *Ind. J. Chem. Technol.* **14** 362-370.
- GUPTA SC, DASS P, SHARMA P and SINGH AV (2002) Removal of ⁵⁸Co, ¹³⁴Cs, and ⁹⁵Zr radioisotopes from aqueous solutions using cellulose iminodiacetic acid chelating cum cation-exchanger. *Desalination* **143** (3) 141-145.
- HARPER TR and KINGHAM NW (1992) Removal of arsenic from waste water using chemical, precipitation methods. *Water Environ. Res.* **64** (3) 200-203.
- HASSAN SSM, MACI SA, BADR IH and ARIDA HA (2000) Novel solid-state ammonium ion potentiometric sensor based on zirconium titanium phosphate ion exchanger. *Anal. Chim. Acta* **427** (1) 21-28.
- HIDALGO M, MASANA A, SALVADO V, MUNOZ M and VALIENTE M (2001) Extraction of palladium with tri-isobutylphosphine sulphide (cyanex 471) in toluene from chloride solutions containing thiocyanate. *Talanta* **38** (5) 483-488.
- HUBICKI Z and LESZCZYNSKA M (2009) Studies of the sorption of palladium (II) ions from model chloride systems onto an ion exchanger containing isothiouraea groups and onto weakly basic anion exchangers of various types. *Adsorpt. Sci. Technol.* **22** (8) 603-614.
- INOUE Y (2003) Studies on the synthetic inorganic ion exchangers-I: Synthesis of stannic phosphate and some of its properties. *Inorg. J. Nucl. Chem.* **26** (12) 2241-2253.
- JEON C and PARK KH (2005) Adsorption and desorption characteristics of mercury (II) ion using aminated chitosan bead. *Water Res.* **39** (16) 3938-3944.
- KANTIPULY C, KATRAGADDA S, CHOW A and GESSER HD (2001) Chelating polymers and related supports for separation and preconcentration of trace metals. *Talanta* **37** (5) 491-517.
- KEPA U, STANCZYK-MAZANEK E and STEPNIAK L (2008) The use of the advanced oxidation process in the ozone + hydrogen peroxide system for the removal of cyanide from water. *Desalination* **233** (1-3) 187-193.
- KHALIL MI and FARAG S (1998) Utilization of some starch derivatives in heavy metal ions removal. *J. Appl. Polym. Sci.* **69** (1) 45-50.
- KHAN AA, INAMUDDIN M and MEZBAUL ALAM M (2004) Preparation, characterization and analytical applications of a new and novel electrically conducting fibrous type polymeric-inorganic composite material: polypyrrole Th(IV) phosphate used as a cation-exchanger and Pb(II) ion-selective membrane electrode. *Mater. Res. Bull.* **40** (2) 289-305.
- KHAN AM, KHAN S and GANAI SA (2009) Role of chemically modified organic resin in selective separation of Hg²⁺ from other metal ions. *Biol. Med.* **1** (2) 127-133.
- KHAN AA and MEZBAUL ALAM M (2003) Synthesis, characterization and analytical applications of a new and novel 'organic-inorganic' composite material as a cation exchanger and Cd (II) ion-selective membrane electrode: polyaniline Sn(IV) tungsto-arsenate. *React. Funct. Polym.* **55** (3) 277-290.
- KIEZYK PR and MACKAY D (2009) Waste water treatment by solvent extraction. *Can. J. Chem. Eng.* **49** (6) 747-752.
- KOZLOWSKI CA, WALKOWIAK TW and KOZIOL JJ (2005) Application of hydrophobic β-cyclodextrin polymer in separation of metal ions by plasticized membranes. *Sep. Purif. Technol.* **46** (3) 136-144.
- LATA AG, GEORGE BK, KANNAN KG and NINAN KN (2003) Synthesis of a polyacrylamide chelating resin and application in metal ion extractions. *J. Appl. Polym. Sci.* **43** (6) 1159-1163.
- MAYA L (2003) Structure and chromatographic applications of crystalline Zr (OPO₃R)₂, R = butyl, lauryl and octylphenyl. *Inorg. Nucl. Chem. Lett.* **15** (5-6) 207-212.
- MONSER L and ADHOUM N (2002) Modified activated carbon for removal of copper, zinc, chromium and cyanide from wastewater. *Sep. Purif. Technol.* **26** (2-3) 137-146.
- NENOV V and BONEV B (2010) Sorption of heavy metals by modified chelating ion exchangers. *Int. J. Environ. Technol. Manage.* **12** (2-4) 343-358.
- OUKI SK and NEUFELD RD (1999) Use of activated carbon for the recovery of chromium from industrial wastewaters. *J. Chem. Technol. Biotechnol.* **70** (1) 3-8.
- PATEL HA, SOMANI RS, BAJAJ HC and JASRA RV (2006) Nanoclays for polymer nanocomposites, paints, inks, greases and cosmetics formulations, drug delivery vehicle and waste water treatment. *Bull. Mater. Sci.* **29** (2) 133-145.
- PEHLIVAN E and ALTUN T (2006) The study of various parameters affecting the ion exchange of Cu²⁺, Zn²⁺, Ni²⁺, Cd²⁺ and Pb²⁺ from aqueous solution on Dowex 50W synthetic resin. *J. Hazard. Mater. B134* 149-156.
- RAMOS RL, RANGEL-MENDEZ JR, MENDOZA-BARRON J, FUENTES-RUBIO L and GUERRERO-CORONADO RM (1997) Adsorption of cadmium from aqueous solution onto activated carbon. *Water Sci. Technol.* **35** (7) 205-211.
- RAO CNR (1963) *Chemical Applications of Infrared Spectroscopy*. Academic Press, New York.
- RENGARAJ S, YEON KH and MOON SH (2001) Removal of chromium from water and wastewater by ion exchange resin. *J. Hazard. Mater.* **87** (1) 273-287.
- SAHA B, IGLESIAS M, DIMMING IW and STREAT M (2000) Sorption of trace heavy metals by thiol containing chelating resins. *Solvent Extr. Ion Exch.* **18** (1) 133-167.
- SANCHEZ JM, HIDALGO M and SALVADO V (2001) The selective adsorption of gold (III) and palladium (II) on new phosphine sulphide-type chelating polymers bearing different spacer arms: Equilibrium and kinetic characterization. *React. Funct. Polym.* **46** (3) 283-291.
- SANTOSA SJ, SISWANTA D, SUDIONO S and SEHOL M (2007) Synthesis and utilization of chitin-humic acid hybrid as sorbent for Cr (III). *Surf. Sci.* **601** (15) 5148-5154.
- SARIN VK, KENT SBH and MERRIFIELD RB (1980) Properties of swollen polymer networks. Solvation and swelling of peptide-containing resins in solid-phase peptide synthesis. *J. Am. Chem. Soc.* **102** (17) 5463-5470.
- SHAH R and DEVI S (1998) Preconcentration and separation of palladium (II) and platinum (IV) on a dithizone anchored poly(vinylpyridine)-based chelating resin. *Anal. Chim. Acta* **341** (2) 217-224.

- SHIM JW, PARK SJ and RYU SK (2001) Effect of modification with HNO₃ and NaOH by pitch- based activated carbon fibers. *Carbon* **39** (11) 1635-1642.
- SINGH AV, GUPTA S and GUPTA SC (1996) Synthesis of a guaran-sulphonic acid cation-exchanger and its application in metal ion removal from underground mine water of the Rajpura Dariba Mines, Udaipur, India. *Desalination* **104** 235-238.
- SINGH A and SARAF SK (2009) Synthesis, characterization and ion-exchanging properties of a novel ion-exchange resin part II. *Int. J. Polym. Mater.* **58** (10) 499-508.
- VESELY V and PEKARK V (2001) Synthetic inorganic ion-exchangers-I: Hydrous oxides and acidic salts of multivalent metals. *Talanta* **19** (3) 219-262.
- WANG CC, CHEN CY and CHANG CY (2002) Synthesis of chelating resins with iminodiacetic acid and its wastewater treatment application. *J. Appl. Polym. Sci.* **84** (7) 1353-1362.
- WAN NAGH WS, ENDUD CS and MAYANAR R (2002) Removal of copper (II) ions from aqueous solution onto chitosan and cross-linked chitosan beads. *React. Funct. Polym.* **50** (2) 181-190.
- YURLOVA L, KRYVORUCHKO A and KORNILOVICH B (2002) Removal of Ni (II) ions from wastewater by micellar-enhanced ultrafiltration. *Desalination* **144** (1) 255-260.
-