



Removal of arsenic from water using copper slag

T. Türk¹, İ. Alp¹, R. Sezer^{2,3}, and C. Arslan³

Affiliation:

¹Department of Mining Engineering, Faculty of Engineering, Karadeniz Technical University, Trabzon, Turkey.

²Department of Metallurgical and Materials Engineering, Faculty of Engineering, Karadeniz Technical University, Trabzon, Turkey.

³Materials Science and Production Technologies Applied Research Center (ATARC), Istanbul Technical University, Istanbul, Turkey.

Correspondence to:

T. Türk

Email:

tugbayilmaz@ktu.edu.tr

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ORCID ID:

T. Türk
<https://orcid.org/0000-0001-5993-2548>

İ. Alp
<https://orcid.org/0000-0002-6032-3528>

R. Sezer
<https://orcid.org/0000-0001-9655-2114>

C. Arslan
<https://orcid.org/0000-0002-0034-7560>

Synopsis

The potential use of copper slag (CS) as an adsorbent for removing arsenic contaminant from water was examined. The influence of solution pH, initial arsenate (As(V)) and arsenite (As(III)) concentrations, and adsorbent dosage was investigated by batch experiments to elucidate the adsorption mechanism of arsenate onto CS. The adsorption kinetics indicated that the second-order kinetic model best described the adsorption process. The adsorption data was analysed by the isotherm models, and was well fitted by the Langmuir model. The maximum removal of As(V) and As(III) achieved was 98.76% and 88.09%, respectively at an adsorbent dose of 10 g/L with an initial As(V) and As(III) concentration of 300 µg/L. This study showed that copper slag is a suitable adsorbent for removal of arsenic from water, with a capacity to reduce arsenic levels to < 10 µg/L, below the limit set for water by World Health Organization.

Keywords

waste management, copper slag, adsorption, arsenic removal.

Introduction

Arsenic is one of the most harmful and toxic elements found in nature. Arsenic is introduced into water through both natural and anthropogenic sources (Hossain, 2006). There have been reports of arsenic poisoning in Argentina, Bangladesh, India, Nepal, Taiwan, the USA, and Turkey (Jean, Bundschuh, and Bhattacharya, 2010). Prolonged exposure to arsenic can lead to serious health problems such as cancers of the liver, kidney, skin, and stomach (Meng, Bang, and Korfiatis, 2000). This suggests an urgent need for efficient and low-cost arsenic removal techniques. The World Health Organization (WHO) has set the drinking water standard at 10 ppb (0.01 mg/L) (WHO, 1993).

Surface and groundwaters contaminated with pollutants can be treated by various methods such as coagulation, precipitation, membrane processes, and adsorption (Mohan, and Pittman, 2007). Many kinds of adsorbents such as iron oxide-loaded slag (Lakshmiathiraj *et al.*, 2006), activated alumina (Thirunavukkarasu, 2002), anionic clays (Türk and Alp, 2012), nano-magnetite (Türk, Alp, and Deveci, 2010) and Fe-hydrotalcite (Türk and Alp, 2010) have been developed to remove arsenic. Previous studies reported that iron-based minerals were useful in the removal of arsenic from water. Although many adsorbents have been developed, few of them are implemented in small settlements to remove pollutants from groundwater (Makris, Sarkar, and Datta, 2006). For these reasons, low-cost waste materials for application in rural areas at a small scale are in high demand. An example of such processes is the use of a waste by-product from the flotation of a copper slag (CS) as a low-cost sorbent. CS is obtained during the smelting and refining of copper and contains copper, zinc, cobalt, and nickel (Gorai, Jana, and Premchand, 2003). During smelting, a copper-rich matte phase (sulphides) and slag phase (oxide-rich) are formed. The slag phase consists predominantly of FeO, Fe₂O₃, and SiO₂, with small amounts of Al₂O₃, CaO and MgO, and Cu, Co, and Ni (as metal) (Davenport *et al.*, 2002).

The annual CS waste production capacity in the Black Sea Copper Works is approximately 150 000t (Alp *et al.*, 2009). As a result, about 1.5–2.0 Mt of CS waste has been disposed of. Improper disposal can result in the toxic metal ions leaching into the groundwater, with ensuing serious environmental problems (Davenport *et al.*, 2002). Due to environmental restrictions and lack of storage space, various options for the treatment of waste materials have been investigated. CS wastes are used in various applications to recycle waste products and to prevent pollution (Alp *et al.*, 2009). The use of CS as an iron source in the production of Portland cement has been reported. Adsorption is an effective and inexpensive method of waste remediation, if adsorbent can be obtained cheaply and in large amounts (Zhang *et al.*, 2004). Arsenic is strongly adsorbed on Fe oxides and hydroxides (Türk and Alp,

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2012) which are abundant in CS. However, a literature review discovered no reports on the use of CS for arsenic removal from water. In this research, the arsenic removal characteristics of CS were investigated.

Experimental

Materials

CS obtained from the Black Sea Copper Works (Samsun, Turkey) was used as adsorbent. The chemical composition of the slag was determined using inductively coupled plasma mass spectrometry (ICP-MS, ACME Analytical Laboratories). The major components in the sample were Fe_2O_3 (59.08%) and SiO_2 (30.60%). The Brunauer-Emmett-Teller method was used to determine the surface area (for CS; 4.81 m^2/g). X-ray diffraction (XRD) patterns of CS were obtained using a RIGAKU D/Max-IIIIC diffractometer. The operating conditions were 35 kV-15 mA, and the sample was scanned between 2° and 60° . A pinch of CS sample was poured onto adhesive carbon tape, and then its surface was covered with a very thin (5–10 μm) gold layer by the physical vapour deposition (PVD) technique. The results of the X-ray analysis and a scanning electron microscopy (SEM, JEOL JSM 5600) secondary ion image of CS are shown in Figure 1. In the example, fayalite (Fe_2SiO_4), magnetite (Fe_3O_4), quartz (SiO_2), and cristobalite (SiO_2) are present.

Adsorption tests

The arsenic stock solution was prepared using sodium arsenate ($\text{Na}_2\text{HAsO}_4 \cdot 7\text{H}_2\text{O}$, from Sigma) and sodium arsenite (NaAsO_2 , Sigma) in pure water. Adsorption tests were carried out in flasks immersed in a temperature-controlled water bath at room temperature (nearly 25°C), and agitated at 130 r/min with a mechanical shaker. The effects of the initial arsenic concentration (10, 50, 100, 300, and 500 $\mu\text{g}/\text{L}$), contact time (1, 3, 5, 7, and 24 hours), adsorbent dosage (1, 5, 10, 20, and 50 g/L), and pH (3, 5, 7, and 9) on the adsorption of arsenic were studied. The isothermal and kinetic studies were conducted using 50 mL of solution (a series of 100 mL flasks) of As(V) and As(III) (10, 50, 100, 300, and 500 $\mu\text{g}/\text{L}$) at an adsorbent dosage of 10 g/L. The collected samples were then centrifuged for 10 minutes at 4000 r/min. Arsenic concentrations were determined by inductively coupled plasma optical emission spectrophotometry (ICP-OES).

The point of zero charge (pH_{pzc}) of CS was found to be 7; this was determined by equilibrating CS with 0.1 M NaCl solutions at different pH values (3–12). Türk, Alp, and Deveci (2010) give details of the procedure.

Results and discussion

Adsorption of arsenic

In adsorption processes, pH is an important parameter. It affects the speciation of metal in aqueous solution and the surface charge of adsorbent (Zhang, and Hideaki, 2005). Figure 2 shows the effect of pH on the adsorption of As(V) and As(III) onto CS. It can be seen that the arsenic adsorbed is independent of pH value. After 5 hours of adsorption, the residual As(V) in solution was determined to be 6.2–2.0 $\mu\text{g}/\text{L}$, and As(III) 55.2–40.0 $\mu\text{g}/\text{L}$ at pH 3–9. The pH_{pzc} value of the CS sample was determined to be pH 7, and at pH values greater than pH_{pzc} , the CS surface is negatively-charged. Although the adsorbent surface and sorbate types were negatively charged, adsorption was not decreased at pH values above pH_{pzc} . This may be due to the buffer effect. In a low-pH medium, the pH of the bulk solution increases in the presence of CS. At high pH values, acid dissociation of CS reduces the pH of the solution (Tian, and Shen, 2009). As a result, a pH value of 9.0 was selected, and all experiments in the study were performed at pH 9.

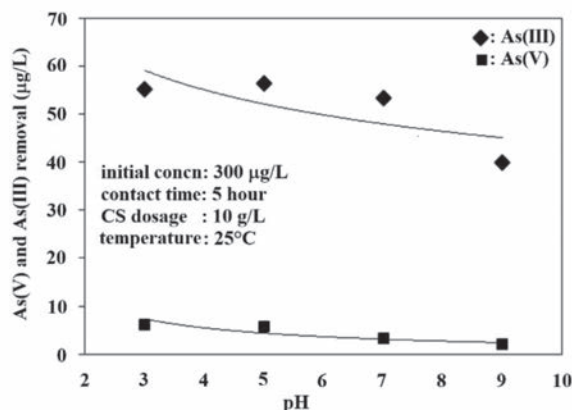


Figure 2—Effect of pH on arsenic removal

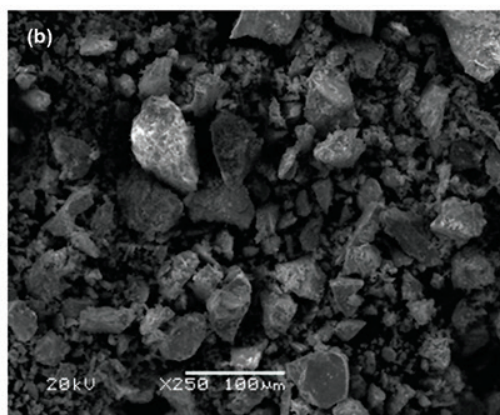
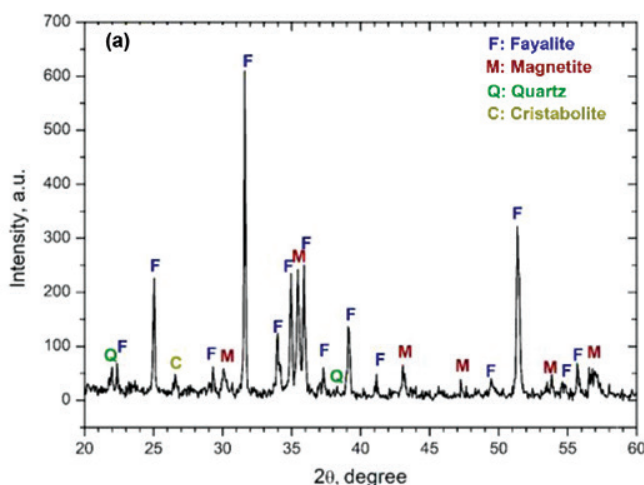


Figure 1—(a) X-ray diffraction spectrum of CS and (b) SEM image

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Figure 3 shows the effect of CS dosage. Arsenic removal efficiency increased with increasing CS. At the end of 5 hours at a CS dosage of 1 g/L, the concentrations of As(V) and As(III) in solution were 104.3 µg/L and 142.3 µg/L respectively, compared to 6.31 and 19.3 µg/L at 10 g/L CS. All further experiments in the study were performed at a dosage of 10 g/L CS.

Kinetic modelling

The removal of As(V) and As(III) versus time is illustrated in Figure 4. CS removed more than 81% of the arsenic within the first hour, and the removal efficiency increased gradually to 99% within 3 hours. The adsorption capacity for arsenic was approximated to 30 µg/g at sorption equilibrium. In these experiments, only 0.5 g of CS was placed in a flask containing up to 100 mL of arsenate solution at a concentration of 300 µg/L. The adsorption kinetics depend strongly on the characteristics of the adsorbent material and adsorbate species, which also influence the adsorption mechanism (Bektaş, Akman, and Kara, 2004). Pseudo-first-order and pseudo-second-order models were used to determine experimental kinetic data models and to understand the mechanism of the adsorption process. The pseudo-first-order equation (Ho, and McKay, 1998) is

$$\ln(q_e - q_t) = \ln q_e - k_1 t \quad [1]$$

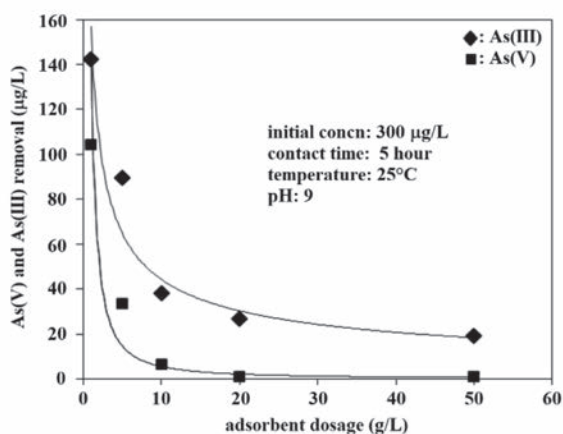


Figure 3—Effect of adsorbent dosage

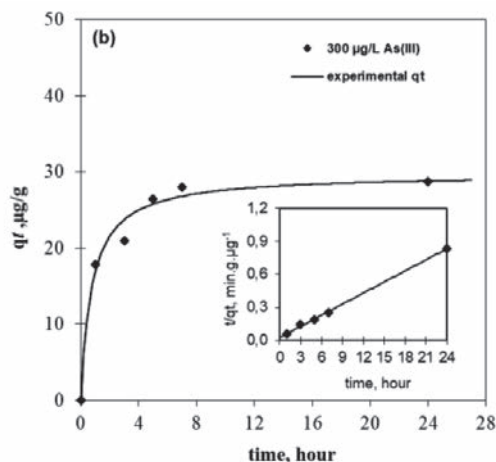
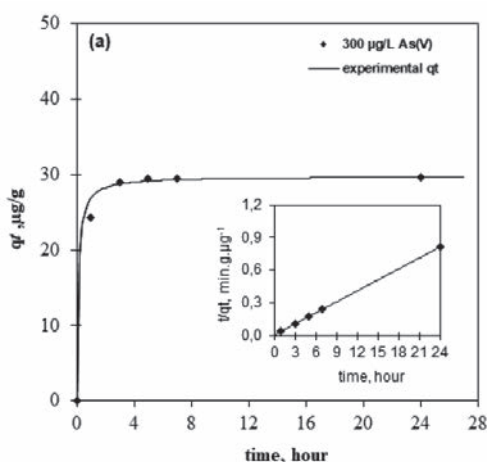


Figure 4—Effect of time on (a) As(V) and (b) As(III) absorption. Initial concn. 300 µg/L; CS dosage 10 g/L; 25°C. The smaller inset graphs show the pseudo-second-order plots

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \quad [2]$$

where q_e is the adsorption capacity (µg/g) at equilibrium, q_t the adsorption capacity (µg/g) at time t , k_1 the pseudo-first-order rate constant (min^{-1}), and k_2 is the pseudo-second-order rate constant ($\text{g mg}^{-1} \text{s}^{-1}$).

The first-order rate constant k_1 and equilibrium adsorption density q_e (Equation [1]) at initial concentrations of As(V) and As(III) were calculated using the slope and intercept of plots of $\log(q_e - q_t)$ versus t . The coefficients of determinations R^2 were found to be 0.947 and 0.854, respectively.

Pseudo-second-order adsorption parameters q_e and k_2 in Equation [2] were determined by plotting t/q_t versus t (Figure 4a, 4b). The coefficient of determination (R^2) of the pseudo-second-order kinetic model is 0.999, higher than the coefficient of determination of the pseudo-first-order model. A high coefficient of determination suggests that the pseudo-second-order kinetic model better represents the adsorption kinetics and that As(V) and As(III) ions are adsorbed on the CS surface through chemical interaction (Bulut, Özacar, and Şengil, 2008). The calculated q_e values also agreed well with the experimental data.

Adsorption equilibrium

The equilibrium adsorption of arsenic as a function of the arsenic concentration was investigated to determine the adsorption capacity of CS. The Langmuir, Freundlich, Dubinin–Radushkevich (D-R), and Temkin models were tested. The adsorption isotherms are shown in Figure 5.

Freundlich isotherm

The Freundlich isotherm is an empirical equation based on a heterogeneous surface and multilayered sorption. The higher the K_f value, the higher the adsorbate affinity (Mohamad, Amir, and Sudabeh, 2013).

The linear form of the equation (Veli, and Akyüz, 2007) is

$$\ln q_e = \ln K_f + \frac{1}{n} \ln C_e \quad [3]$$

where q_e is the amount of adsorbate at equilibrium (µg/g), C_e is the equilibrium concentration of adsorbate (µg/L), n is adsorption intensity, and K_f is the Freundlich constant. The linear plot of \ln

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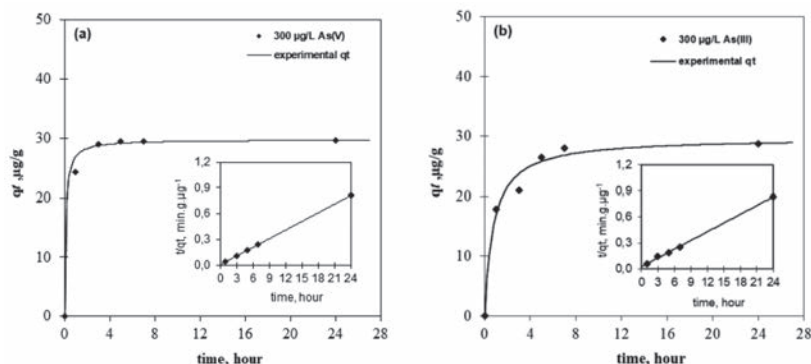


Figure 5—Equilibrium isotherms for adsorption

q_e and $\ln C_e$ gives a slope of $1/n$ and intercept of $\ln K_f$. In theory, values of n from 1 to 10 indicate that the adsorbate is adsorbed appropriately on the adsorbent (Konicki *et al.*, 2012). In this study, $n = 1.73$. The Freundlich adsorption isotherm model is shown in Figures 6 and 7. The coefficient of determination of this model is lower than those for the other three isotherms (Table I).

Langmuir isotherm

The Langmuir isotherm is the most widely used isotherm for removal of contaminants from solutions (Crini *et al.*, 2009; Özacar and Şengil, 2005). The Langmuir isotherm suggests that adsorption occurs in a single layer on the surface of a homogeneous adsorbent (Foo and Hameed, 2010).

The linear form of the conventional Langmuir model (Kundu and Gupta, 2007) is

$$\frac{1}{q_e} = \frac{1}{bQ C_e} + \frac{1}{Q} \quad [4]$$

where C_e is the equilibrium concentration ($\mu\text{g/L}$), b is the adsorption constant ($\text{L}/\mu\text{g}$), q_e is the amount of arsenic at equilibrium ($\mu\text{g/g}$), and Q is the adsorption capacity ($\mu\text{g/g}$).

The Langmuir constants and coefficient of determination (R^2) values were obtained from the slope and intercept of the linear plot of $1/q_e$ versus $1/C_e$ (Figures 6 and 7, Table I). The Langmuir model was best fitted to experimental data with high R^2 values.

The Langmuir isotherm can be explained by the R_L parameter (Kundu, and Gupta, 2007), which is defined by Equation [5]

$$R_L = \frac{1}{1 + bC_0} \quad [5]$$

where C_0 ($\mu\text{g/L}$) is the initial concentration and b ($\text{L}/\mu\text{g}$) is the Langmuir constant. R_L values may indicate unfavourable ($R_L > 1$), linear ($R_L = 1$), favourable ($0 < R_L < 1$), or irreversible ($R_L = 0$) adsorption (Tekin, 2006). The values of R_L for our experiments were found to be < 0.04 , indicating that the adsorption by CS is favourable.

The Langmuir isotherm is well adapted to the experimental data. It is an indication of single layer adsorption of arsenic on CS (Bulut, Özacar, and Şengil, 2008). The Langmuir isotherm model is widely used in explaining chemical adsorption (Dubinin, 1960). The isotherm study was conducted using different arsenic concentrations (50–300 $\mu\text{g/L}$) at 25°C, pH 9, 0.5 g dosage of CS, and adsorption time 5 hours. For CS, the maximum adsorption capacities in arsenate and arsenite solution were found to be 109 and 84.8 $\mu\text{g/g}$ respectively, using the Langmuir isotherm. The maximum adsorption capacities of some iron-based adsorbent materials are compared with this value. Thirunavukkarasu, Viraraghavan, and Subramanian (2003) studied the adsorption of

As(V). on iron oxide-coated sand. They found that the adsorption capacity was 45 $\mu\text{g/g}$. Guo, Stüben, and Berner (2007) studied As(V) removal using haematite. The adsorption capacity of haematite was found to be 204 $\mu\text{g/g}$ (As(V) from a 1000 $\mu\text{g/L}$ solution at an adsorbent dosage of 10 g/L). Türk and Alp (2010) reported adsorption capacities of 80 $\mu\text{g/g}$ for As(III) on Fe-hydroxalcite.

Dubinin-Radushkevich (D-R) isotherm

The Dubinin-Radushkevich (D-R) isotherm model is used to distinguish the physical and chemical adsorption of arsenic ions on CS (Dubinin, 1960). This model is written as (Kundu, and Gupta, 2007)

$$q_e = Q_D \exp(-B_D \varepsilon^2) \quad [6]$$

and linearized as

$$\ln q_e = \ln Q_D - B_D \varepsilon^2 \quad [7]$$

where Q_D is the maximum capacity (mol/g), B_D is the D-R constant (mol^2/kJ^2), and ε is the Polanyi potential and is equal to $\varepsilon = RT \ln(1 + \frac{1}{C_e})$.

Figures 6 and 7 show the plot of $\ln q_e$ against ε^2 . The coefficient of determination $R^2 = 0.9166$. B_D was found to be $0.34 \text{ mol}^2/\text{kJ}^2$, and Q_D was 35.63 mol/g .

The free energy of adsorption (E) was determined using the relationship $E = \frac{1}{\sqrt{2B_D}}$ (Kundu, and Gupta, 2007). The value of E is significant when ascertaining the type of adsorption. If the E value is less than 8 kJ/mol, adsorption is physical, and if it is 8–16 kJ/mol, adsorption is chemical (Islam, Mishra, and Patel, 2011). In this investigation the value of E was found to be 0.84 kJ/mol (Table I), thus in this case, the adsorption is physical. Similar results have been reported for the adsorption of arsenic by iron-oxide-coated granular activated charcoal (Ananta, Saumen, and Vijay, 2015).

Temkin isotherm

The Temkin isotherm is based on adsorbate-adsorbent interactions. This model ignores deficient and extremely high concentrations, and the free energy of adsorption is a function of the surface coating. The Temkin isotherm is given as (Mall *et al.*, 2005):

$$q_e = \frac{RT}{b} \ln(AC_e) \quad [8]$$

$$\frac{RT}{b} = B \quad [9]$$

The linear form of the Temkin isotherm is

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$$q_e = B \ln A + B \ln C_e \quad [10]$$

where A is the equilibrium binding constant ($L/\mu g$), b is the Temkin constant (J/mol), R is the gas constant ($8.314 J/mol K$), and T is the absolute temperature (K) equivalent. The adsorption data was analysed and the values of the Temkin constants A and B and coefficient of determination established; these are listed in Table I, and the theoretical plot of this isotherm is shown in Figures 6 and 7. According to the coefficients of determination (R^2), the order of suitability of the models was Langmuir, Freundlich, D-R, and Temkin (Figures 6, 7 and Table I). Thus, the Langmuir model is the most suitable for describing single-layer adsorption of arsenic on the CS surface. The adsorption capacities (Q) for As(V) and As(III) in the Langmuir model were 109.460 and 84.83 $\mu g/g$, respectively. As a result, the isothermic order is Langmuir > Temkin > Freundlich, > D-R isotherm.

Conclusion

In this study, the removal of As(V) and As(III) from aqueous solution by copper slag (CS) was demonstrated. The pH effect

Table I

Calculated Isotherm parameters (10 g/L CS, 25°C)

Parameter	As(V)		As(III)	
	Value	R ²	Value	R ²
Langmuir				
Q ($\mu g/g$)	109.29	0.991	84.83	0.981
b	0.083		0.030	
Freundlich				
K_f	8.88	0.895	1.37	0.917
n	1.73		1.00	
D-R				
Q_D	35.63	0.916	26.80	0.816
B_D	0.34		0.23	
Temkin				
A	2.64	0.982	0.09	0.950
B	11.4		15.5	

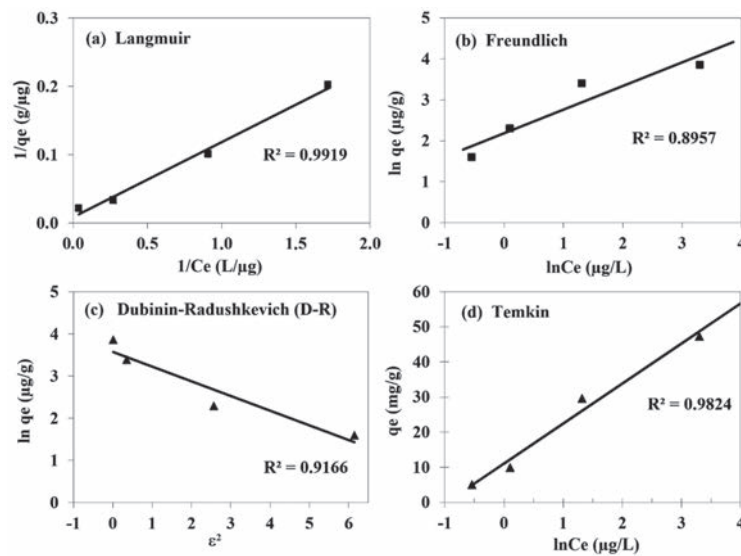


Figure 6—Adsorption isotherm plots for As(V)

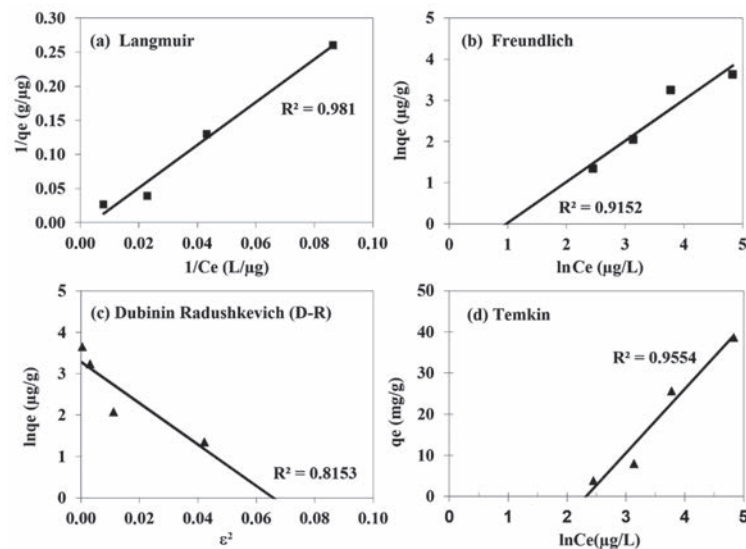


Figure 7—Adsorption isotherm plots for As(III)

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study showed that the amount of arsenic adsorbed by CS was independent of initial pH value. Kinetic studies showed that the rate of adsorption accords with the pseudo-second-order kinetic model and that arsenic ions are adsorbed on the CS surface through chemical interaction. The adsorption data was analysed by Langmuir, Freundlich, Dubinin-Radushkevich (D-R) and Temkin isotherm models and the data was well fitted using the Langmuir isotherm model. The Langmuir isotherm model is widely used in explaining chemical adsorption. The maximum adsorption capacities of the sorbent in As(V) and As(III) solution were found to be 109.29 and 84.83 µg/g, respectively. The RL value (< 0.04) showed that CS is favourable for the removal of arsenic from aqueous solutions. Under optimal conditions (adsorbent dosage 10 g/L, solution pH 9, temperature 25°C, and arsenic concentration of 300 µg/L the adsorbent was able to remove 98.76% of the arsenate and 85.43% of the arsenite. CS can be used as a practical, easy to use, and inexpensive adsorbent for arsenate and arsenite removal from water.

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References

- ALP, I., DEVECI, H., YAZICI, E.Y., TÜRK, T., and SUNGUN, H. 2009. Potential use of pyrite cinders as raw material in cement production: Results of industrial scale trial operations. *Journal of Hazardous. Materials*, vol. 166. pp. 144-149. doi:10.1016/j.jhazmat.2008.10.129
- ANANTA, S., SAUMEN, B., and VIJAY, V. 2015. Adsorption isotherm, thermodynamic and kinetic study of arsenic (III) on iron oxide coated granular activated charcoal. *International Research Journal of Environmental Sciences*, vol. 4. pp. 64-77.
- Bektaş, N., Akman A.B., and Kara, S. 2004. Kinetic and equilibrium studies in removing lead ions from aqueous solutions by natural sepiolite. *Journal of Hazardous. Materials*, vol. 12, pp. 115-122. doi:10.1016/j.jhazmat.2004.04.015
- BULUT, E., ÖZACAR, M., and ŞENGLİ, İ.A. Equilibrium and kinetic data and process design for adsorption of congo red onto bentonite. *Journal of Hazardous. Materials*, vol. 154. pp. 613-622. doi:10.1016/j.jhazmat.2007.10.071
- CRINI, G., PEINDY, H.N., GIMBERT, F., and ROBERT, C. 2007. Removal of C.I. Basic Green 4, (Malachite Green) from aqueous solutions by adsorption using cyclodextrin-based adsorbent: kinetic and equilibrium studies. *Separation and Purification Technology*, vol. 53. pp. 97-110. doi:10.1016/j.seppur.2006.06.018
- DAVENPORT, W.G., KING, M., SCHLESINGER, M., and BISWAS, A.K. Extractive Metallurgy of Copper. 4th edn. Pergamon Press, Oxford.
- DUBININ, M.M. 1960. The potential theory of adsorption of gases and vapors for adsorbents with energetically nonuniform surface. *Chemical Reviews*, vol. 60. pp. 235-266. doi: 10.1021/cr60204a006
- FOO, K.Y. and HAMEED, B.H. 2010. Insights into the modeling of adsorption isotherm systems. *Chemical Engineering Journal*, vol. 156. pp. 2-10. doi:10.1016/j.cej.2009.09.013
- GORAI, B., JANA, R.K., and PREMCHAND. 2003. Characteristics and utilization of copper slag - A review. *Resource Conservation and Recycling*, vol. 39. pp. 299-313. doi:10.1016/S0921-3449(02)00171-4
- GUO, H., STÜBEN, D., and BERNER, Z. 2007. Removal of arsenic from aqueous solution by natural siderite and hematite. *Applied Geochemistry*, vol. 22. pp. 1039-1051. doi:10.1016/j.apgeochem.2007.01.004
- HO, Y.S. and MCKAY, G. 1998. Kinetic model for lead(II) sorption onto peat. *Adsorption Science and Technology*, vol. 16. pp. 243-255. doi:10.1177/026361749801600401
- HOSSAIN, M.F. 2006. Arsenic contamination in Bangladesh - An overview. *Agriculture, Ecosystems & Environment*, vol. 113. pp. 1-16. doi:10.1016/j.agee.2005.08.034
- ISLAM, M., MISHRA, P.C., and PATEL, R. 2011. Fluoride adsorption from aqueous solution by a hybrid thorium phosphate composite. *Chemical Engineering Journal*, vol. 166. pp. 978-985. doi:10.1016/j.cej.2010.11.087
- JEAN, J.S., BUNDSCHUH, J., and BHATTACHARYA, P. 2010. Arsenic in Geosphere and Human Diseases. Balkema, CRC Press.
- KONICKI, W., PELECH, I., MIJOWSKA, E., and JASIŃSKA, I. Adsorption of anionic dye Direct Red 23 onto magnetic multi-walled carbon nanotubes-Fe₃C nanocomposite: Kinetics, equilibrium and thermodynamics. *Chemical Engineering Journal*, vol. 210. pp. 87-95. doi:10.1016/j.cej.2012.08.025
- KUNDU, S. and GUPTA, A.K. 2007. Adsorption characteristics of As (III) from aqueous solution on iron oxide coated cement (IOCC). *Journal of Hazardous. Materials*, vol. 142. pp. 97-104. doi:10.1016/j.jhazmat.2006.07.059
- LAKSHMIPATHIRAJ, P., NARASIMHAN, B.R.V., PRABHAKAR, S., and BHASKAR RAJU, S.G. 2006. Adsorption of arsenate on synthetic goethite from aqueous solutions. *Journal of Hazardous. Materials*, vol. 136. pp. 281-287. doi:10.1016/j.jhazmat.2005.12.015
- MAKRIS, K.C., SARKAR, D., and DATTA, R. 2006. Evaluating a drinking-water waste by-product as a novel sorbent for arsenic. *Chemosphere*, vol. 64. pp. 730-741. doi:10.1016/j.chemosphere.2005.11.054
- MALL, I.D., SRIVASTAVA, V.C., AGARWALL, N.K., and MISHRA, I.M. 2005. Removal of Congo Red from aqueous solution by bagasse fly ash and activated carbon. Kinetic study and equilibrium isotherm analysis. *Chemosphere*, vol. 61. pp. 492-501. doi:10.1016/j.chemosphere.2005.03.065
- MENG, X.G., BANG, S., and KORFIATIS, G.P. 2000. Effects of silicate, sulfate, and carbonate on arsenic removal by ferric chloride. ii, vol. 34. pp. 1255-1261. DOI:10.1016/S0043-1354(99)00272-9
- MOHAMAD, A.B., AMIR, H.M., and SUDABEH, P. 2013. Thermodynamic analysis of reactive Red 198 removal from synthetic wastewater by using multiwall carbon nanotubes. *Health Scope*, vol. 2. pp. 149-155. doi: 10.17795/jhealthscope-13438
- MOHAN, D. and PITTMAN, C.U. 2007. Arsenic removal from water/wastewater using adsorbents - A critical review. *Journal of Hazardous. Materials*, vol. 142. pp. 1-53. doi: 10.1016/j.jhazmat.2007.01.006
- Özacar, M. and Şengil, İ.A. 2005. Adsorption of metal complex dyes from aqueous solutions by pine sawdust. *Bioresource Technology*, vol. 96. pp. 791-795. doi:10.1016/j.biortech.2004.07.011
- TERKIN, N., KADINCI, E., DEMIRBAŞ, Ö., ALKAN, M., and KARA, A. 2006. Adsorption of polyvinylimidazole onto kaolinite. *Journal of Colloids and Interface Science*, vol. 296. pp. 472-479. doi:10.1016/j.jcis.2005.09.049
- THIRUNAVUKKARASU, O.S. 2001. Removal of arsenic from drinking water, PhD thesis, Regina University, Regina, Saskatchewan.
- THIRUNAVUKKARASU, O.S., VIRARAGHAVAN, T., and SUBRAMANIAN, K.S. 2003. Arsenic removal from drinking water by iron-oxide coated sand. *Water, Air, & Soil Pollution*, vol. 142. pp. 95-111. doi:10.1023/A:1022073721853
- TIAN, C.S. and SHEN, Y.R. 2009. Structure and charging of hydrophobic material/water interfaces studied by phase-sensitive sum-frequency vibrational spectroscopy. *Proceedings of the National Academy of Sciences*, vol. 106. pp. 15148-15153. doi:10.1073/pnas.0901480106
- TÜRK, T. and ALP, İ. 2010. Adsorption of As(III) from water using Mg-Fe-hydroxalate (FeHT). *Ekoloji*, vol. 74. pp. 77-88 [in Turkish]. doi: 10.5053/ekoloji.2010.7411
- TÜRK, T. and ALP, İ. 2012. Adsorption of arsenic from borated water using Mg-Fe-hydroxalate (FeHT). *Ekoloji*, vol. 84. pp. 98-106. [in Turkish]. doi: 10.5053/ekoloji.2012.8411
- TÜRK, T., ALP, İ., and DEVECI, H. 2010. Adsorptive removal of arsenite from water using nanomagnetite. *Desalination and Water. Treatment*, vol. 24. pp. 302-307. doi:10.5004/dwt.2010.1404
- VELI, S. and AKYÜZ, B. 2007. Adsorption of copper and zinc from aqueous solutions by using natural clay. *Journal of Hazardous. Materials*, vol. 149. pp. 226-233. doi:10.1016/j.jhazmat.2007.04.109
- WORLD HEALTH ORGANIZATION. 1993. WHO guidelines for drinking water quality. Geneva. 41 pp.
- ZHANG, F. and HIDEAKI, I. 2005. Iron oxide-loaded slag for arsenic removal from aqueous system. *Chemosphere*, vol. 60. pp. 319-325. doi:j.chemosphere.2004.12.019
- ZHANG, W., SINGH, P., PALING, E., and DELIDES, S. 2004. Arsenic removal from contaminated water by natural iron ores. *Minerals Engineering*, vol. 17. pp. 517-524. doi:j.mineng.2003.11.020