



A nonlinear prediction model, incorporating mass transfer theory and expert rules, for refining low-carbon ferrochrome

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Synopsis

We present an optimal oxygen-blowing system with expert rules to improve the efficiency of refining low-carbon ferrochrome. A nonlinear model based on mass transfer theory, the principles of heat transfer, and the principles of high-temperature chemical reactions for refining low-carbon ferrochrome are established. The model is mainly used to control the oxygen supply rate during argon-oxygen top-bottom double-blown refining, thereby controlling the refining temperature and reducing the carbon content. Twenty production tests using a 5 t argon-oxygen refining furnace demonstrate the effectiveness of the system and reliability of the nonlinear model. A comparison of the model data with the experimental data shows that although the model fails to predict the silicon content in the final refined product, it can predict the contents of the main components at the refining end-point and the refining temperature accurately.

Keywords

prediction model, end-point control, mass transfer theory, expert rules.

Introduction

Low-carbon ferrochrome (LCFeCr) is used for trimming additions in stainless steel production as well as other specialist applications. LCFeCr can be produced by various processes, such as the low-basicity furnace process, Perrin process, electric-silicothermic process, bottom-blown converter process, argon-oxygen decarburization (AOD) process, and vacuum furnace reactor process. (Weitz and Garbers-Craig, 2016; Heikkinen and Fabritius, 2012; Booyesen *et al.*, 1998; Bhonde, Ghodgaonkar, and Angal, 2007).

The AOD process has been extensively researched and numerous models have been developed. Noting the change in the partial pressure of carbon monoxide with the bath height, Roy and Robertson (1978a, 1978b) introduced the heat balance to take into account the non-isothermal nature of the bath and proposed a mathematical model for stainless steelmaking. On the basis of the mass and heat balances, Reichel and Szekely (1995) developed a mathematical model of the AOD process for stainless steel. Görnerup and Sjöberg (1999) mathematically modelled the AOD/Creusot-Loire Uddeholm process. Wei and Zhu (2002a, 2002b) modelled the whole AOD process, including the top, side, and combined blowing operations in the decarburization phase and the reduction practice.

In these models, the main product is generally stainless steel. In recent years, many experts have studied the production of LCFeCr by the AOD process (Ville-Valtteri *et al.*, 2013a, 2013b; Qiu *et al.*, 2013; Deo and Srivastava, 2003). In the refining process, the oxygen supply rate, refining duration, refining temperature, and other factors will have a great impact on the composition of the products.

The existing models for stainless steel are not suitable for LCFeCr refining of due to the differences in composition and technology (Heikkinen and Fabritius, 2012). Industry experts and researchers have attached importance to the establishment of mathematical models of LCFeCr production technology, and have conducted research on various aspects of the process. Some scholars have established corresponding mathematical models. Booyesen *et al.* (1998) developed a bottom-blown AOD model based on phenomenological heat and kinetic mass transfer. Bhonde, Ghodgaonkar, and Angal (2007) developed a mathematical model according to the metallothermic reduction, triplex process, and vacuum technique. Akimov *et al.* (2013) developed a mathematical model based on the results of industrial smelts and thermodynamic analysis. Although some models describing ferrochrome production have been developed, most of the models focus on individual phenomena rather than the total process. Some important parameters, such as refining temperature, carbon oxidation rate, and mass transfer

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rate (Wei and Zhu, 2002a; Ville-Valtteri *et al.*, 2017), have not been reliably determined to date. There is still no accurate mathematical model for LCFerCr production by the top-bottom double-blown AOD process. Many producers still rely on the experience of workers in the refining of LCFerCr. There is therefore an urgent need to establish an accurate model to describe this process.

Mathematical model for low-carbon ferrochrome

Modelling conditions

An AOD furnace with a capacity of 5 t was selected for the experimental work. Oxygen is blown into the furnace through a top lance, and is mixed with the ferrochrome melt. The main lance and the secondary lance are installed at the bottom of the furnace. The main lance blows oxygen and argon. The secondary lance blows argon only.

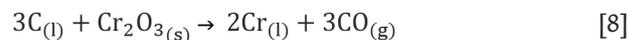
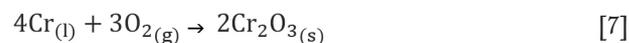
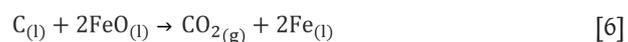
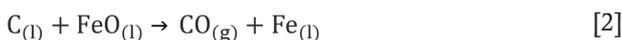
The solid raw materials are melted in the electric arc furnace, then the molten ferrochrome is poured into the argon-oxygen top-bottom double-blown refining furnace for refining. The initial temperature in the refining furnace is about 1600K. The electric arc furnace provides feed material of fixed composition for the refining furnace. The carbon content is 6.5–8.5%, the silicon content 1.5–2.0%, the chromium content 57–65%, the iron content 25–35%, and the other elements are less than 0.1%.

The following assumptions are made in order to simplify the modeling of the refining process with a small refining capacity:

- (1) During the refining process, elements such as Fe, Cr, C, and Si in the high-temperature melt are oxidized. The oxidation reactions and the reduction reactions take place continuously.
- (2) It is assumed that the redox reaction in the furnace is uniform during the refining process.
- (3) Others trace elements in the melt are ignored.
- (4) The refining temperature is controlled between 1800 and 2100K. The temperature inside the furnace is assumed to be uniform.

Chemical reactions in the refining process

The reactions during argon-oxygen refining of LCFerCr occur at normal pressure, in the presence of sufficient oxygen, and at high temperature. The chemical reactions are complex. The main chemical reactions that occur during the stabilization period can be described by Equations [1]–[8] (Bi, 1975; Cai, 1995; Wei *et al.*, 2011).



where (l) indicates a component of the melt, (s) a component of the slag, and (g) a component of the gaseous phase.

The refining process entails the dynamic equilibrium of high-temperature chemical reactions. The gas produced by the chemical reaction is discharged as flue gas. The molten oxide matter produced by the chemical reactions become the slag. Based on the above reactions, a mathematical model is established to predict the melt content during the refining process and the reaction rates, which can be used to guide production.

Mathematical model of the refining process

Mathematical model based on heat balance

The ultimate goal of refining LCFerCr is to remove carbon and other impurities. Elemental oxidation occurs at the gas-liquid interface (Manuela, Sergey, and Masamichi, 1997), and the elements at the interface are transported by mass transfer (Olga, Genii, and Pavel, 2014). Therefore, the oxidation rate of an element can be calculated from its mass transfer rate (Ville-Valtteri *et al.*, 2017). In this way, the elemental composition of the ferrochrome melt can be predicted.

Heat balance calculation

The heat generated and heat lost during the refining process, are made up of several components (Table I).

According to the law of energy conservation and heat analysis, the heat balance equation is obtained as:

$$\begin{aligned} & W_m c_{p,m} T_0 + W_s c_{p,s} T_0 + \\ & \left(\frac{dO_2}{dt} \rho_{O_2} c_{p,O_2} + \frac{dAr}{dt} \rho_{Ar} c_{p,Ar} \right) T_c + \\ & \frac{W_m}{100} \left(\frac{dW[C]\%}{dt} \Delta H_C + \frac{dW[Cr]\%}{dt} \Delta H_{Cr} + \right. \\ & \left. \frac{dW[Si]\%}{dt} \Delta H_{Si} \right) dt = W_m c_{p,m} T \\ & \left[1 + \frac{dt}{100} \left(\frac{dW[C]\%}{dt} + \frac{dW[Cr]\%}{dt} + \frac{dW[Si]\%}{dt} \right) \right] \\ & + c_y \frac{2p_y V_y M_y}{RT} T + Q_t + Q_d \end{aligned} \quad [9]$$

Table I

Heat input items and output items

Heat input	Heat output
Enthalpy of molten ferrochrome fed into the furnace	Enthalpy of ferrochrome melt at the end of refining
Enthalpy released by the oxidation reactions	Enthalpy of slag
Heat of the gas	Heat loss due to flue gas
Heat of additives	The loss heat by the furnace body

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The symbols used in this and subsequent equations are defined at the end of the paper.

The heating rate of the furnace is given by Equation [10]. It can be seen that the rate of temperature increase is mainly affected by the oxidation reactions and the heat loss.

$$\begin{aligned} \frac{dT}{dt} = & \left[\left(\frac{dW[C\%]}{dt} \Delta H_C + \frac{dW[Cr\%]}{dt} \Delta H_{Cr} + \right. \right. \\ & \left. \left. \frac{dW[Si\%]}{dt} \Delta H_{Si} \right) - c_{p,m} T \frac{dW[C\%]}{dt} + \right. \\ & \left. \frac{dW[Cr\%]}{dt} + \frac{dW[Si\%]}{dt} \right] \frac{W_m + W_s}{100c_m W_m + 100c_s W_s} \\ & - c_y \frac{2p_y q_y M}{RT} T - (q_t + q_d) \end{aligned} \quad [10]$$

Oxidation rate calculation

Diffusion of carbon through the gas-liquid boundary layer has been shown to be the rate-controlling mechanism (He *et al.*, 1994). According to mass transfer theory, the mass transfer rate of carbon across the gas-liquid interface can be calculated as (Guo, 2006):

$$\frac{dn_C}{dt} = Ak_d (c_{[C]} - c_{[C]s}) \quad [11]$$

where $A = 4\pi r^2$ is the bubble surface area; k_d is mass transfer coefficient of oxygen; $c_{[C]}$ is the concentration of carbon in molten ferrochrome, and $c_{[C]s}$ is the concentration of carbon at the bubble interface of molten ferrochrome.

$$c_{[C]} - c_{[C]s} = c_{[O]} - c_{[O]s} = \Delta c_{[O]} \quad [12]$$

where $c_{[O]}$ is the concentration of oxygen in molten ferrochrome in mol/m³; $c_{[O]s}$ is the concentration of oxygen at the bubble interface in mol/m³; and $\Delta c_{[O]}$ is the equilibrium concentration of oxygen (Guo, 2006).

$$\Delta c_{[O]} = \frac{\rho}{Ar_0} \Delta \omega_{[O]} \quad [13]$$

where r_0 is the initial radius of the bubble (6×10^{-5} m) (Guo, 2006), and $\Delta \omega_{[O]}$ is the degree of supersaturation. The reported values for the parameter $\Delta \omega_{[O]}$ range from 0.015% to 0.025% (Guo, 2006). In this work, $\Delta \omega_{[O]} = 0.02\%$.

$$k_d = 2 \sqrt{\frac{D_C}{\pi t_e}} = 4 \left(\frac{g D_C}{9\pi^2 r} \right)^{\frac{1}{4}} \quad [14]$$

where D_C is the diffusion coefficient of carbon in molten ferrochrome (5×10^{-9} m²/s) (Li, 2004); t_e is element contact time; g is acceleration due to gravity (9.81 m/s²); and r is the bubble radius.

$$r = \left\{ \frac{14RT}{4\sqrt{g}} \left(\frac{3D}{\pi} \right)^{\frac{1}{2}} \left(\frac{\Delta c_{[O]}}{\rho g} \right) \left[\ln \left(h + \frac{p_g}{\rho g} \right) - \ln \frac{p_g}{\rho g} \right] \right\}^{\frac{4}{7}} \quad [15]$$

where R is the gas constant (-8.314 J/(mol·K)); ρ is the density of molten ferrochrome (7×10^3 kg/m³), p_g is standard atmosphere, and h is the depth of the furnace.

Substituting R , ρ , p_g , h and Equation [13], the bubble radius is calculated as follows:

$$r = 0.5 \times 10^{-5} T \quad [16]$$

Substituting Equations [13], [14], and [16] into Equation [11] yields:

$$\frac{dn_C}{dt} = 1.13 \times 10^{-8} T^{\frac{7}{4}} \quad [17]$$

$dW[C]\%$ is the rate of decline in carbon content in the molten ferrochrome, which is a function of the rate of ascent of bubbles through the melt. The relationship between dn_C and $dW[C]\%$ is as follows (Guo, 2006):

$$dW[C]\% = \frac{M_C dn_C}{1000W} \times 100 = \frac{12 \times 10^{-4} dn_C}{W} \quad [18]$$

The rate of change in carbon content in molten ferrochrome can be expressed as follows:

$$\begin{aligned} \frac{dW[C]\%}{dt} &= \frac{M_C dn_C}{1000W dt} \times 100 = \\ \frac{12 \times 10^{-4} dn_C}{W dt} &= \frac{1.36 \times 10^{-11} T^{\frac{7}{4}}}{W} \end{aligned} \quad [19]$$

From the temperature, refining time, and the weight of molten ferrochrome, the oxidation rate of carbon in the furnace can be predicted. Thereby the carbon content in the melt can be also calculated.

The oxidation rate of chromium at the gas-liquid interface is positively correlated with the mass transfer rate of chromium while the temperature is higher than 1873K. The mass transfer rate can be described as follows:

$$\begin{aligned} \frac{dn_{Cr}}{dt} &= Ak_{Cr} (c_{[Cr]} - c_{[Cr]s}) \\ &= 0.41 \times 10^{-8} T^{\frac{7}{4}} \end{aligned} \quad [20]$$

where $k_{Cr} = 4 \left(\frac{g D_{Cr}}{9\pi^2 r} \right)^{\frac{1}{4}}$ is the mass transfer coefficient of

chromium, and $D_{Cr} = 1.8 \times 10^{-9}$ m²/s is the diffusion coefficient of chromium in molten ferrochrome.

The oxidation rate of chromium, which can be used to predict the chromium content in the alloy, is given as follows:

$$\frac{dW[Cr]\%}{dt} = \frac{M_{Cr} dn_{Cr}}{1000W dt} \times 100 = \frac{2.13 \times 10^{-11} T^{\frac{7}{4}}}{W} \quad [21]$$

From the temperature, refining time, and the weight of molten ferrochrome, the oxidation rate of chromium in the furnace can be predicted. Thereby the chromium content in the melt can be also calculated.

The oxidation of silicon is similar to that of carbon. The mass transfer rate of silicon can be described as follows:

$$\frac{dn_{Si}}{dt} = Ak_{Si} (c_{[Si]} - c_{[Si]s}) = 0.1 \times 10^{-8} T^{\frac{7}{4}} \quad [22]$$

where $k_{Si} = 4 \left(\frac{g D_{Si}}{9\pi^2 r} \right)^{\frac{1}{4}}$ is the mass transfer coefficient of

silicon, and D_{Si} is the diffusion coefficient of silicon in molten ferrochrome (0.44×10^{-9} m²/s).

The oxidation rate of silicon is given by Equation [23].

$$\frac{dW[Si]\%}{dt} = \frac{M_{Si} dn_{Si}}{1000W dt} \times 100 = \frac{1.23 \times 10^{-11} T^{\frac{7}{4}}}{W} \quad [23]$$

The oxidation rate of silicon in the furnace can be predicted from the temperature, refining time, and the weight of molten ferrochrome. Thereby the silicon content in the bath can be calculated.

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The standard enthalpies of formation of the respective oxides (at 298K) involved in the following equations and the relevant heat capacities at constant pressure with the enthalpies of solution formation were all taken from Wei and Zhu (2002a) and Turkdogan (1980).

$$\Delta H_c = \Delta H_{CO} - \Delta H_c - \frac{\Delta H_{O_2}}{2} \quad \Delta H_c = 11852 - (2.367T_g + 1.708 \times 10^{-4}T_g^2 + 3.853 \times 10^3/T_g) \quad [24]$$

$$\Delta H_{Cr} = \Delta H_{Cr_2O_3} - 2\Delta H_{Cr} - 3\frac{\Delta H_{O_2}}{2} \quad \Delta H_{Cr} = 11519 - (1.148T + 4.4 \times 10^{-5}T^2 + 1.5 \times 10^4/T) \quad [25]$$

$$\Delta H_{Si} = \Delta H_{SiO_2} - \Delta H_{Si} - \Delta H_{O_2} \quad \Delta H_{Si} = 30658 - (2.15T + 1.45 \times 10^{-4}T^2) \quad [26]$$

Heat dissipation

Heat dissipation calculation for furnace body

An elevation diagram of the argon-oxygen refining furnace is shown in Figure 1. The furnace is composed of two parts: a furnace body and a furnace cap. The middle part of the furnace body is cylindrical; the upper and lower parts are truncated cones. The ratio between the depth of the molten bath and the height of the furnace chamber is 1:3. The ratio between the molten bath depth and the diameter of the molten pool surface is 1:2. The furnace body is regarded as a cylinder for convenience of calculation. The refractory lining of the shaft has two layers. The inner (working) layer consists of chrome-magnesia refractory brick with a thickness of 350 mm. The outer layer is insulated and reinforced with refractory clay bricks. Its thickness is 115 mm.

In the LCFer production process in an AOD furnace, the temperature in the furnace is about 1873–2073K, the outer wall temperature is about 623K, and the ambient temperature is 293–303K. In the refining process, the heat losses consist of the conduction heat and radiant heat from the furnace body, the conduction heat and radiant heat at the bottom of the furnace, and the heat loss in the flue gas (Farrera-Buenrostro *et al.*, 2019). According to the equation for the heat transfer rate, the heat loss from the furnace body can be calculated.

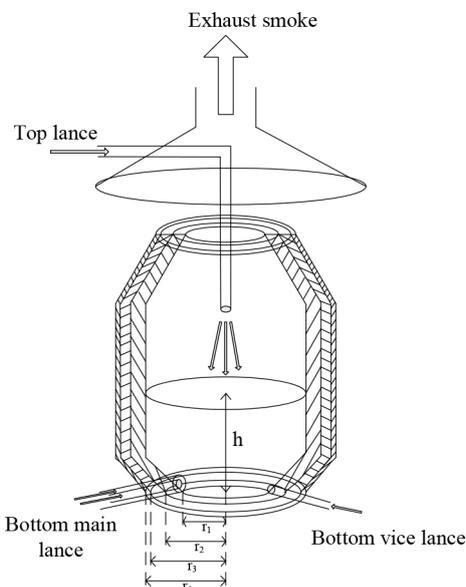


Figure 1—Schematic diagram of an AOD furnace

The exothermic reactions occur within a few minutes, and no heat is transferred to the lining during this time. It is assumed that each part of the furnace lining has the same heat transfer characteristics, which is convenient for theoretical analysis and modelling. Therefore, the heat transfer state is assumed to be stable. The heat transfer rate of a three-layer lining furnace body (Liao, *et al.* 2017) can be described as follows:

$$q_t = \frac{2\pi L(T_1 - T_4)}{\frac{1}{\lambda_1} \ln \frac{r_2}{r_1} + \frac{1}{\lambda_2} \ln \frac{r_3}{r_2} + \frac{1}{\lambda_3} \ln \frac{r_4}{r_3}} \quad [27]$$

where T_1 is the temperature in the furnace, T_4 is the outer wall temperature, λ_1 is the thermal conductivity of magnesia brick, λ_2 is the thermal conductivity of magnesia-chrome brick, λ_3 is the thermal conductivity of the shell, and r_i is the radius. In this work, $\lambda_1 = 3.49$, $\lambda_2 = 1.42$, $\lambda_3 = 60$, $r_1 = 0.7$ m, $r_2 = 1.0$ m, $r_3 = 1.06$ m, and $r_4 = 1.15$ m. Therefore, the heat loss from the body of the furnace, q_t , is $41.87\Delta T$.

The heat transfer rate at the bottom of the furnace (Liao, *et al.* 2017) can be described as follows:

$$q_d = \lambda A \frac{dT}{dn} = \frac{A(T_1 - T_4)}{\frac{\delta_1}{\lambda_1} + \frac{\delta_2}{\lambda_2} + \frac{\delta_3}{\lambda_3}} \quad [28]$$

where $\delta_1 = 0.6$ m is the thickness of the magnesia brick, $\delta_2 = 0.12$ m is the thickness of magnesia-chrome brick, and $\delta_3 = 0.1$ m is the thickness of shell. Therefore, the heat loss from the bottom of the furnace, q_d , is $3.12\Delta T$.

Flue gas heat loss calculation with expert rules

The flue gas is composed mainly of carbon monoxide, oxygen, argon, and some dust. Dust, which consists of solid particles, can be collected by the dust collector and the heat loss calculated together with that of the slag. Therefore, the flue gas is considered to be composed of carbon monoxide, carbon dioxide, oxygen, and argon only.

The flue gas heat can be calculated as follows:

$$Q_y = CmT_g \quad [29]$$

where C_y is the specific heat capacity, m is the material quality, and T_g is the temperature of the flue gas.

The gas quality can be calculated as follows, assuming it follows the Ideal Gas Law:

$$m = pVM/RT_g \quad [30]$$

where p is the gas pressure, V is the volume of gas, M is the molar mass of gas, and T_g is the gas temperature.

If the flue gas flow is substituted for the gas volume, the heat loss rates via the flue gas can be obtained as follows:

$$q_y = C_y \frac{2P_y \frac{dQ_y}{dt} M_y}{RT_g} T_g \quad [31]$$

where C_y is the specific heat capacity of the flue gas, P_y is the flue gas pressure, $\frac{dQ_y}{dt}$ is the flue gas flow, M_y is the molar mass of

flue gas, and R is the gas constant. From the oxygen utilization efficiency, the oxygen partial ratio, and the specific heat capacity of the gas, the specific heat capacity of the flue gas can be calculated. The pressure of the flue gas is equal to atmospheric pressure. From the oxygen utilization efficiency, the oxygen

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Table II

Flue gas me parameters

$P_y = 1.01325 \times 10^5 \text{Pa}$	$\text{CO}_2 = 0.9184$
$C_{Ar} = 0.5238$	$C_{\text{CO}} = 1.0447$
$\eta = 0.8$	$\text{MO}_2 = 32 \times 10^{-3} \text{kg/mol}$
$M_{Ar} = 40 \times 10^{-3} \text{kg/mol}$	$M_{\text{CO}} = 28 \times 10^{-3} \text{kg/mol}$

partial ratio, and the molar mass of the gas, the molar mass of the flue gas can be calculated. The flue gas temperature is about half of the furnace temperature (He, 2016). The values of some flue gas parameters (Chen, 1984; Gorges, Pulvemacher, and Rubens, 1978) are shown in Table II:

In order to lessen the oxidation of chromium and avoid excessive temperatures, the ratio of oxygen to argon needs to be adjusted with the carbon content and refining temperature of the furnace (Andersson *et al.*, 2012). The physical and chemical reactions in the high-temperature refining process are very complex. The rate of reaction will be affected by factors such as the flow rate of oxygen, the amount of feed, and the rate of diffusion of elements in the bath (Odenthal, 2010). Therefore, the amounts of argon and oxygen should be adjusted according to expert rules. In the early stage of refining, slag-forming agent should be added to prevent chromium volatilization. In the later stage of refining, deoxidizer should be added to remove oxygen in the bath. Chromium ore should be added to increase the chromium content of the end product. Figure 2 shows the flow chart of the model strategy. The whole process is divided into three stages based on expert rules. The corresponding information is listed in Table III.

From the ratio of the gases and the oxygen utilization efficiency, the flow rate of flue gas for the three stages can be calculated separately.

For the first stage:

$$\frac{dQ_y}{dt} = (1 - 85\%) \frac{dO_2}{dt} + \frac{1}{6} \frac{dO_2}{dt} + 30\% \times 85\% \frac{dO_2}{dt} = 0.58 \frac{dO_2}{dt} \quad [32]$$

For the second stage:

$$\frac{dQ_y}{dt} = (1 - 65\%) \frac{dO_2}{dt} + \frac{3}{7} \frac{dO_2}{dt} + 80\% \times 65\% \frac{dO_2}{dt} = 1.3 \frac{dO_2}{dt} \quad [33]$$

For the third stage:

$$\frac{dQ_y}{dt} = (1 - 35\%) \frac{dO_2}{dt} + \frac{4}{5} \frac{dO_2}{dt} + 35\% \frac{dO_2}{dt} = 1.8 \frac{dO_2}{dt} \quad [34]$$

The specific heat capacity of the flue gas can be calculated based on the ratio of carbon monoxide, oxygen, and argon as follows:

For the first stage:

$$C_y = \frac{(1-85\%)}{(1-85\%)+1/6+30\% \times 85\%} c_{O_2} + \frac{1/6}{(1-85\%)+1/6+30\% \times 85\%} c_{Ar} + \frac{30\% \times 85\%}{(1-85\%)+1/6+30\% \times 85\%} c_{CO} = 0.26c_{O_2} + 0.29c_{Ar} + 0.45c_{CO} = 0.86 \quad [35]$$

For the second stage:

$$C_y = \frac{(1-65\%)}{(1-65\%)+3/7+80\% \times 65\%} c_{O_2} + \frac{3/7}{(1-65\%)+3/7+80\% \times 65\%} c_{Ar} + \frac{80\% \times 65\%}{(1-65\%)+3/7+80\% \times 65\%} c_{CO} = 0.27c_{O_2} + 0.33c_{Ar} + 0.40c_{CO} = 0.84 \quad [36]$$

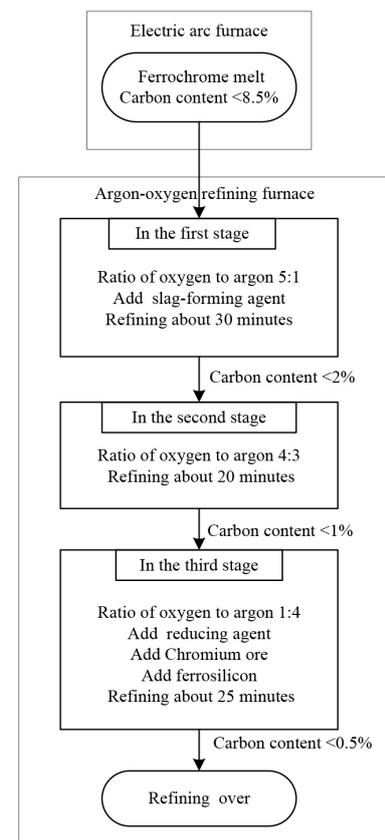


Figure 2—Flow chart of model strategy

Table III

The data of three stages

	The volume ratio of oxygen to argon	Oxygen utilization efficiency(%)	The oxygen partial ratio
The first stage	5:1	80–90	$C \leq 30\%$, $Si \leq 30\%$
The second stage	4:3	60–70	$C \leq 80\%$, $Si \leq 10\%$
The third stage	1:4	30–40	$C \leq 90\%$, $Si \leq 0\%$

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For the third stage:

$$C_y = \frac{(1-35\%)}{(1-35\%)+4/5+35\%} c_{O_2} + \frac{4/5}{(1-35\%)+4/5+35\%} c_{Ar} + \frac{35\%}{(1-35\%)+4/5+35\%} c_{CO} = 0.33c_{O_2} + 0.24c_{Ar} + 0.20c_{CO} = 0.77 \quad [37]$$

The molar mass of the flue gas can be calculated based on the carbon monoxide, oxygen, and argon contents as follows:

For the first stage:

$$M_y = \frac{(1-85\%)}{(1-85\%)+1/6+30\% \times 85\%} M_{O_2} + \frac{1/6}{(1-85\%)+1/6+30\% \times 85\%} M_{Ar} + \frac{85\%}{(1-85\%)+1/6+30\% \times 85\%} M_{CO} = 0.26M_{O_2} + 0.29M_{Ar} + 0.45M_{CO} = 32.52 \times 10^{-3} \text{ kg mol}^{-1} \quad [38]$$

For the second stage:

$$M_y = \frac{(1-65\%)}{(1-65\%)+3/7+80\% \times 65\%} M_{O_2} + \frac{3/7}{(1-65\%)+3/7+80\% \times 65\%} M_{Ar} + \frac{65\%}{(1-65\%)+3/7+80\% \times 65\%} M_{CO} = 0.27M_{O_2} + 0.33M_{Ar} + 0.40M_{CO} = 33.04 \times 10^{-3} \text{ kg mol}^{-1} \quad [39]$$

For the third stage:

$$M_y = \frac{(1-35\%)}{(1-35\%)+4/5+35\%} M_{O_2} + \frac{4/5}{(1-35\%)+4/5+35\%} M_{Ar} + \frac{35\%}{(1-35\%)+4/5+35\%} M_{CO} = 0.36M_{O_2} + 0.45M_{Ar} + 0.19M_{CO} = 34.84 \times 10^{-3} \text{ kg mol}^{-1} \quad [40]$$

The heat loss rates *via* flue gas in the three stages are calculated as follows:

For the first stage:

$$q_y = C_y \frac{2P_y \frac{dQ_y}{dt} M_y}{RT} T = 3.99 \times 10^3 \frac{dO_2}{dt} \quad [41]$$

For the second stage:

$$q_y = C_y \frac{2P_y \frac{dQ_y}{dt} M_y}{RT} T = 8.99 \times 10^3 \frac{dO_2}{dt} \quad [42]$$

For the third stage:

$$q_y = C_y \frac{2P_y \frac{dQ_y}{dt} M_y}{RT} T = 12.04 \times 10^3 \frac{dO_2}{dt} \quad [43]$$

It can be shown from the model that the bath temperature is related to the oxygen blowing rate, the oxygen blowing time, and the weight of the ferrochrome. Hence the bath temperature can be predicted as follows.

For the first stage:

$$\frac{dO_2}{dt} = \frac{1 \times 10^{-14} T^7}{w} \left(1.96 \times 10^4 - 1.69T - 0.83 \times 10^{-4} T^2 - \frac{1.07 \times 10^4}{T} \right) - 1.13 \times 10^{-2} T + 7.02 - 0.25 \times 10^{-3} \frac{dT}{dt} \quad [44]$$

For the second stage:

$$\frac{dO_2}{dt} = \frac{1 \times 10^{-14} T^7}{w} \left(0.87 \times 10^4 - 0.75T - 0.37 \times 10^{-4} T^2 - \frac{0.47 \times 10^4}{T} \right) - 0.5 \times 10^{-2} T + 3.11 - 0.11 \times 10^{-3} \frac{dT}{dt} \quad [45]$$

For the third stage:

$$\frac{dO_2}{dt} = \frac{1 \times 10^{-14} T^7}{w} \left(0.65 \times 10^4 - 0.56T - 0.27 \times 10^{-4} T^2 - \frac{0.35 \times 10^4}{T} \right) - 0.37 \times 10^{-2} T + 2.33 - 0.08 \times 10^{-3} \frac{dT}{dt} \quad [46]$$

Experimental results and analysis

Twenty production tests were carried out in a 5 t argon-oxygen refining furnace, and the results compared with the predicted values. As shown in Figures 3–6, the changes in carbon, silicon, and chromium content and temperature during the refining process can be clearly seen.

The reducing agent is fed into the refining process in the third stage and affects the contents of C and Si in the reaction. Therefore, the actual contents of C and Si in the third stage show a large deviation from the predicted values.

Temperature is the most important parameter in the whole argon-oxygen refining process. The control of C, Si, and Cr contents and temperature are closely related. The changes in C, Si, and Cr content are shown in Figures 3–5, and the corresponding temperature change trend is shown in Figure 6.

In the first reaction stage, the top lance supplies oxygen from the top of the furnace, the main lance supplies the oxygen-argon mixture from the bottom of the furnace, and the secondary lance supplies argon from the bottom of the furnace. The ratio of oxygen to argon is 5:1. Elements such as carbon, chromium, and silicon in the ferrochrome melt are oxidized rapidly, and the refining temperature increases rapidly due to the exothermic oxidation reactions. In order to reduce the volatilization of chromium, 150 kg of slag-forming agent is added, and the refining temperature drops slightly.

At the end of the first reaction stage, the temperature reaches 2047K, the carbon content falls to 1.24%, and the silicon content is 0.49%. This takes about 30 minutes, and the total amount of oxygen blown is about 105 Nm³. The oxidation rate of chromium slows down. In the second reaction stage, the top lance stops supplying oxygen, the main lance supplies the oxygen-argon mixture, and the secondary lance supplies argon. The ratio of oxygen to argon is 4:3. The total supply of oxygen is reduced, the supply of argon is increased, and the temperature of the molten pool is allowed to rise in a controlled fashion. The oxidation of

A nonlinear prediction model, incorporating mass transfer theory and expert rules

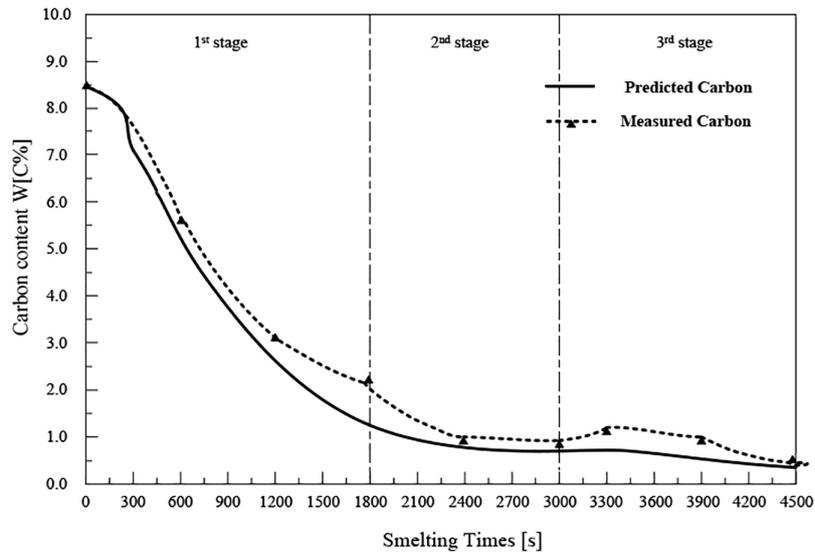


Figure 3—Change in carbon content during the refining process

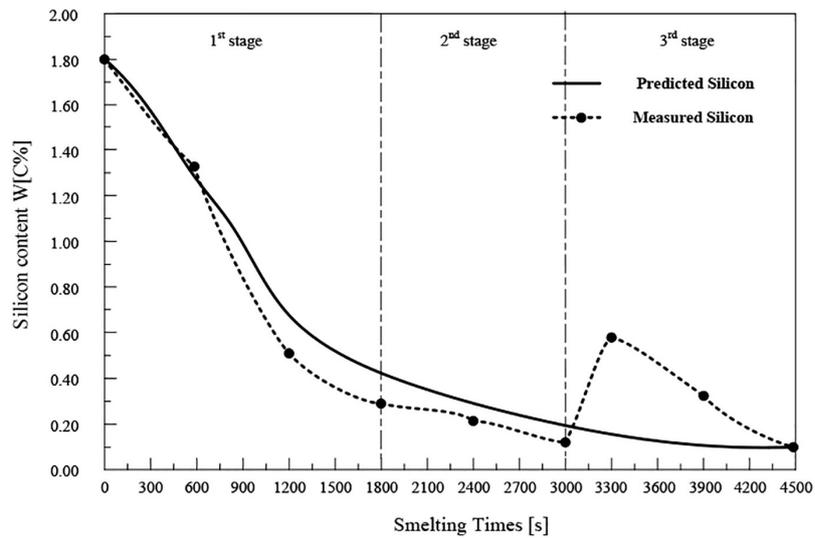


Figure 4—Change in silicon content during the refining process

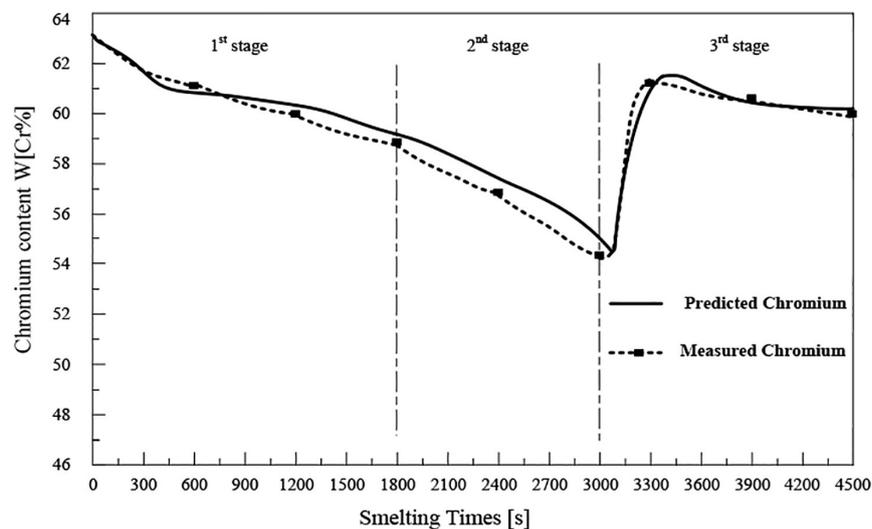


Figure 5—Change in chromium content during the refining process

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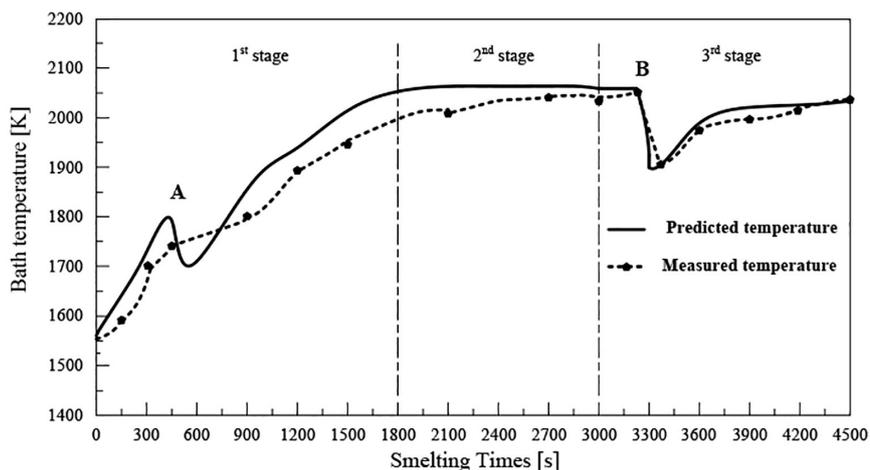


Figure 6—Temperature change during the refining process

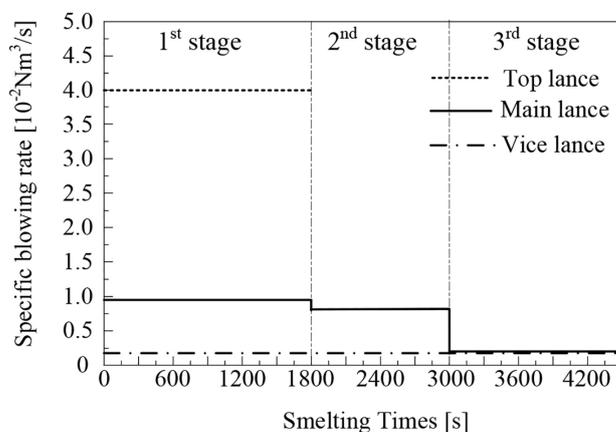


Figure 7—Specific blowing rate

carbon and oxygen takes place mainly in the molten pool, the carbon content is further reduced, the oxidative heat release and heat dissipation are balanced, and the temperature of the bath tends to be stable. After 20 minutes of refining, the total amount of oxygen blown is about 55 m³, the carbon content is 0.96%, and the temperature is 2053K. Entering the third reaction stage, the ratio of oxygen to argon blown is 1:4. 100 kg of reducing agent is added to reduce the chromium from the slag. In order to increase the chromium content at the end of the refining process, 200 kg of chromium ore is added, and 150 kg of ferrosilicon is added to remove oxygen. After 25 minutes, refining is stopped. The total oxygen supply is about 30 m³, the final carbon content is 0.43%, the chromium content is 60.1%, and the temperature is 2036K.

The specific blowing rate is shown in Figure 7.

The test data is shown in Table IV, which lists the liquid phase mass, initial temperature, composition, and refining time.

The predicted and experiments data for the 20 production tests are listed in Table V. The maximum absolute errors between the model predictions and the measured values of carbon content, chromium content, and temperature are 0.07 mass%, 2.92 mass%, and 6K, respectively. The average relative errors between the predicted and the measured values of carbon content, chromium content, and temperature are 5.49%, 1.79%, and 0.21%, respectively. The results show that the model can accurately predict the contents of the main components at the

refining end point and the refining temperature, which proves that the model is reliable. However, the predicted value of the silicon content deviates considerably from the measured value. The reason for this is that deoxidizers were been added in the later stages of refining.

Conclusion

A mathematical model has been established for refining low-carbon ferrochrome based on high-temperature oxidation reactions, mass transfer theory, and expert rules. The effectiveness of the model was verified by comparing the predicted compositions of the metal phase with the measured values from 20 test runs in a 5 t argon-oxygen refining furnace. Statistical analysis of the results showed that the prediction accuracy is high in terms of final carbon content, chromium content, and metal bath temperature. The model can accurately estimate the refining end point composition and the temperature, and can be used to guide the production of low-carbon ferrochrome by the AOD process. The experimental data, such as furnace capacity, geometry, and additives, can be adjusted to the actual production requirements.

List of symbols

W_m	Mass of liquid ferrochrome (kg)
W_s	Mass of slag (kg)
T_g	Temperature of flue gas (K)

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t_c	Contact time (s)
g	Standard gravity (m/s^2)
$\Delta\omega_{[1]}$	Degree of supersaturation (mol/m^3)
C_y	Specific heat capacity of the flue gas at constant pressure ($J/(kg\cdot K)$)
V_y	Flue gas volume (m^3)
p_y	Flue gas pressure (Pa)
q_y	Heat loss rate via flue gas (J/s)
k_i	Mass transfer coefficient of element i (m/s)
D_i	Diffusion coefficient (m^2/s)
r^i	Radius (m)
h	Depth of the furnace (m)
q_t	Heat transfer rate through furnace body (J/s)
q_d	Heat transfer rate through the bottom of the furnace (J/s)
λ_i	Heat conductivity ($W/(m\cdot K)$)
δi	Thickness of the diffusion boundary layer (m)
ρ_i	Density of element i (kg/m^3)
π	Mathematical constant (m)

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